lonic Bonding

Name: Hour:

Information: Polyatomic Ions

The word, "polyatomic" means "many atoms". A polyatomic ion, therefore, is an ion that is made of more than one atom. An example of a polyatomic ion is the sulfate ion, $SO_4{}^5$. Sulfate is composed of one sulfur atom and four oxygen atoms and overall sulfate has a negative two charge.

Some polyatomic ions:

Sulfate: SO₄²⁻ Phosphate: PO₄²⁻ Nitrate: NO₅²⁻ Cyanide: CN²⁻ Ammonium: NH₄²⁻ Chlorate: ClO₅²⁻ Acetate: C₂H₃O₂²⁻ Hydroxide: OH²⁻ Carbonate: CO₃²⁻

Critical Thinking Questions

- 1. What do all of the polyatomic ions that have the suffix "-ate" have in common?
- 2. Which two atoms do you think compose the polyatomic ion called "silicate"?
- 3. What is the difference between calcium nitride and calcium nitrate?

Information: Writing Formulas With Polyatomic Ions

First of all, you must remember that you can never change the formula for a polyatomic ion. Sulfate is <u>always</u> SO_4 and never S_2O_8 or something else. Following are some examples of chemical formulas that contain polyatomic ions.

Ammonium chloride is formed from one ammonium ion (NH₄*) and one chloride ion (Cl') to give the formula: NH₄Cl. Sodium sulfate requires two sodium ions (Na*) because sulfate (SO₄*) has a negative two charge; the formula is: Na₂SO₄.

Consider calcium hydroxide. Calcium has a positive two charge (Ca²⁺) and hydroxide has a negative one charge (OH'). We need two hydroxide ions to combine with one calcium ion so that the overall charge ends up being zero. We write calcium hydroxide like Ca(OH)₂.

Following are some more examples: potassium acetate: KC₂H₃O₂ magnesium nitrate: Mg(NO₃)₂ barium phosphate: Ba₃(PO₄)₂ calcium carbonate: CaCO₃

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