

Unit 5: Acids & Bases

Practice Set 2: 2-1 to 2-4 K_a , K_b , K_w and pH

Answers

1. Calculate $[H^+]$ in a 2.00 L solution of hydrogen chloride in which 3.65 g of HCl is dissolved. K_a for HCl is very large.

We will need to know the concentration of the acid, HCl.

Concentration is measured as molarity which has the units $\frac{mol}{L}$

Since the question tells us the mass of HCl in grams, we will need to convert grams to moles. To do so we need to know the molar mass of HCl.

element	Molar Mass (g/mol)	$\times$	No. Atoms		
H	1.0	$\times$	1	=	1.0
Cl	35.5	$\times$	1	=	35.5
Molar Mass compound (g/mol)				=	36.5 $\frac{g}{mol}$

Unit analysis may be used to calculate the concentration of the acid solution:

$$[HCl] = 3.65g \times \frac{mol}{36.5g} \times \frac{1}{2.00L} = 0.05M$$

Next write a balanced equation: $HCl_{(aq)} \rightarrow H^+_{(aq)} + Cl^-_{(aq)}$

Since HCl is a strong acid, it breaks down completely into ions. BECAUSE OF THIS, since there is a 1:1 relationship between HCl and H^+ , the $[HCl]$ will equal the $[H^+]$:

$$[H^+] = [HCl] = 0.05 M$$

Answer

2. Calculate $[H^+]$ in a solution containing 3.20 g of HNO_3 in 250 mL of solution. Nitric acid is a very strong acid.

This is a similar question – again we need to find $[H^+]$ for a strong acid. Begin by calculating the concentration of the acid solution.

This will involve converting grams HNO_3 into moles HNO_3 :

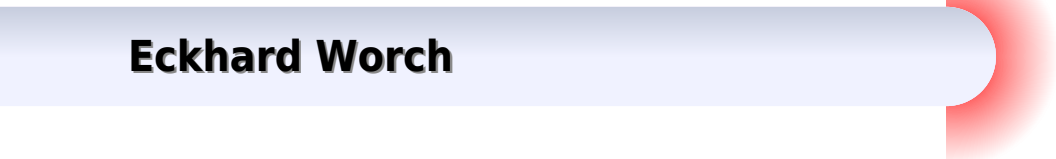
element	Molar Mass (g/mol)	$\times$	No. Atoms		
H	1.0	$\times$	1	=	1.0
N	14.0	$\times$	1	=	14.0
O	16.0	$\times$	3	=	48.0
Molar Mass compound (g/mol)				=	63.0 $\frac{g}{mol}$

Unit analysis may be used to calculate the concentration of the acid solution:

$$[HNO_3] = 3.20g \times \frac{mol}{63.0g} \times \frac{1}{0.250L} = 0.203M$$

Describing Acids And Bases Answers

Eckhard Worch



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