

Chemical Reactions



REACTIONS

11.1 Describing Chemical Reactions



Essential Understanding Chemical reactions are represented by balanced chemical equations.

Lesson Summary

Introduction to Chemical Equations. A chemical equation uses symbols, and sometimes words, to show the reactants and products of a chemical reaction.

- A skeleton equation uses chemical formulas to represent reactants and products, but it does not indicate the relative amounts of each.
- Anything that enters into a reaction, such as heat or a catalyst, but is not a reactant or product is shown above or below the yields arrow in the equation.

Balancing Chemical Equations Chemical equations are balanced to show that mass is conserved during chemical reactions.

- A balanced equation shows the relative amounts of reactants and products, and it contains equal numbers of each type of atom on both sides of the equation.
- Chemical equations are balanced by using coefficients in front of the chemical formulas for the reactants and the products in a skeleton equation.



BUILD Math Skills

Boloncing Equations All chemical equations must be balanced because of the law of conservation of mass, which states that matter cannot be created or destroyed. So, the number of atoms that you start with at the beginning of the reaction must equal the number of atoms that you end up with.

For example, the reaction $2\text{Mg} + O_s \rightarrow 2\text{MgO}$ follows the law of conservation of mass because you start with 2 magnesium atoms and 2 oxygen atoms and you end up with 2 magnesium atoms and 2 oxygen atoms. You can think of the \rightarrow as an = sign.



Turn the page to learn more about balancing equations.

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- 1. Understanding the eBook 1describing Chemical Reactions Section Review Answer Key
 - The Rise of Digital Reading 1describing Chemical Reactions Section Review Answer Key
 - Advantages of eBooks Over Traditional Books
- 2. Identifying 1describing Chemical Reactions Section Review Answer Key
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
- 3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an 1describing Chemical Reactions Section Review Answer Key
 - User-Friendly Interface
- 4. Exploring eBook Recommendations from 1describing Chemical Reactions Section Review Answer Key
 - Personalized Recommendations
 - 1describing Chemical Reactions Section Review Answer Key User Reviews and Ratings
 - 1describing Chemical Reactions Section Review Answer Key and Bestseller Lists
- 5. Accessing 1describing Chemical Reactions Section Review Answer Key Free and Paid eBooks

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 - Adjustable Fonts and Text Sizes of 1describing Chemical Reactions Section Review Answer Key
 - Highlighting and Note-Taking 1describing Chemical Reactions Section Review Answer Key
 - Interactive Elements 1describing Chemical Reactions Section Review Answer Key
- 8. Staying Engaged with 1describing Chemical Reactions Section Review Answer Key
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers 1describing Chemical Reactions Section Review Answer Key
- 9. Balancing eBooks and Physical Books 1describing Chemical Reactions Section Review Answer Key
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection 1describing Chemical Reactions Section Review Answer Key
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine 1describing Chemical Reactions Section Review Answer Key
 - Setting Reading Goals 1describing Chemical Reactions Section Review Answer Key
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of 1describing Chemical Reactions Section Review Answer Key
 - Fact-Checking eBook Content of 1describing Chemical Reactions Section Review Answer Key
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development

- Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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