

11.1

DESCRIBING CHEMICAL REACTIONS

Section Review

Objectives

- 1. Represent the processes of a reaction.
- 2. Explain how to write balanced equations.
- 3. List the steps for writing complete ionic equations.

Vocabulary

- 1. chemical equation
- 2. balanced equation
- 3. catalyst
- 4. coefficient
- 5. skeletal equation

Part A Completion

Write the coefficient for each in these chemical equations, and balance and label every chemical reaction. Indicate the physical states of the reactants and products.

1. A chemical reaction is usually represented by a chemical _____ 2. _____
 The chemical formula for a chemical compound is _____ 3. _____ The mass
 number for an element is represented by the _____ 4. _____ An element's atomic
 mass is represented by _____ 5. _____ A chemical equation is written with an
 arrow pointing to the right. The arrow is _____ 6. _____ or _____ 7. _____
 indicates that the right-hand side of the equation is _____ 8. _____
 The equation for a chemical reaction is balanced by changing the _____ 9. _____ of
 the chemical formulas in the equation.

Chemical equations are written with the physical states of the reactants and products.
 The physical state of a substance is _____ 10. _____ that is, a solid _____ 11. _____, a liquid _____ 12. _____, a gas _____ 13. _____, or
 dissolved in water _____ 14. _____
 In a chemical equation, the physical states of the reactants and products are indicated by
 the letters _____ 15. _____, _____ 16. _____, _____ 17. _____, _____ 18. _____, or
 _____ 19. _____.

- 1. equation
- 2. formula
- 3. atomic mass
- 4. mass
- 5. coefficient
- 6. arrow
- 7. right-pointing arrow
- 8. product
- 9. coefficient
- 10. state
- 11. solid
- 12. liquid
- 13. gas
- 14. aqueous
- 15. (s)
- 16. (l)
- 17. (g)
- 18. (aq)
- 19. catalyst

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