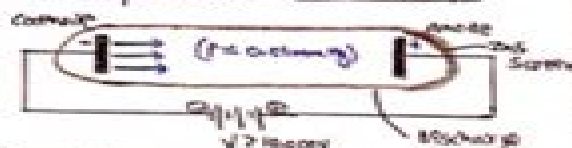


Discovery of Fundamental Particles:

Electron :-

- Discovered by J.J. Thomson in 'Cathode Ray Experiment'.
- When high voltage ($>10^4$ V) is applied across a discharge tube containing a gas at very low pressure (0.001 mm), then a ray flows from cathode to anode.

→ This ray is called 'Cathode Ray'.



Properties of Cathode Rays :-

- 1) Travels in straight line with high speed (fast stream).
- 2) Deviates like particle in motion.
- 3) Carry charge (gets deflected towards the plate in electric field).
- 4) Carry momentum (deviates like particle).
- 5) Affects photographic plate and produces 'Green Glow' when strike wall beyond anode.
- 6) Can pass through metal sheets.
- 7) Composed of very charged electrons.

- Path of e^- in uniform EF is parabolic.
- Path of e^- in uniform mag. field is circular ($r = \frac{mv}{eB}$).
- Charge by mass ratio ($\frac{e}{m}$) = 1.76×10^{11} C/kg.
- Charge of e^- was discovered by Millikan oil drop.
- Resting mass of e^- (m_e) = $\frac{h}{\lambda} = \frac{h}{c \cdot \lambda}$ (from Planck's eqn).

- Characteristics of cathode rays are independent of gas nature of gas taken in discharge tube.

Proton :-

- Discovered by Goldstein & Rutherford in 'Cathode Ray Experiment'.
- Experimental setup is same as above. The only diff. is that here instead of anode, cathode is perforated (cathode with holes).

→ 'Canal/Anode Ray' flows from holes of cathode.



Atomic Structure - 2

Properties of Canal Rays :-

- 1) Travel in str. line with very high speed.
- 2) Carry +ve charge & momentum.
- 3) Behave like particle in motion.
- 4) Characteristics of canal rays depends upon nature of gas taken in discharge tube.
- 5) If the gas taken in discharge tube, then ray consists of +vely charged particles called 'proton'.

$$\text{Mass of proton} = 1.6726 \times 10^{-27} \text{ kg}$$

Neutron :-

- Discovered by J. Chadwick.
- When Beryllium (Be) nucleus bombarded with α -particles (He nuclei) then neutrons emitted.



Particle	Representation	Charge	Mass
α -particle	${}^4_2\text{He}$	+2e	4 amu
β -particle	e^-	-e	$\frac{1}{1840}$ amu
Proton	${}^1_1\text{H}$ or p^+	+e	1.007 amu
Neutron	${}^1_0\text{n}$	zero	1.008 amu

Specific Charge = $\frac{\text{Charge of } \alpha\text{-particle}}{\text{Mass of } \alpha\text{-particle}}$

Order of SC: $e^- > p > \alpha\text{-particle} > n$

Isotopes :-

- 1) Isotopes → Same Z but different A.
Eq: ${}^1_1\text{H}$, ${}^2_1\text{H}$, ${}^3_1\text{H}$
- 2) Isobars → Same A but different Z.
Eq: ${}^4_2\text{He}$, ${}^4_1\text{H}$
- 3) Isodiaphers → Same no. of (n-p) or (A-2Z)
Eq: ${}^{14}_6\text{C}$, ${}^{14}_7\text{N}$
- 4) Isotones → Same no. of neutrons (A-Z)
Eq: ${}^{14}_6\text{C}$, ${}^{15}_7\text{N}$
- 5) Isodiaphers → Species having same number of e^-
Eq: Na^+ , Mg^{2+} , Al^{3+} , C^{4+} , N^{5+} , O^{6+} , F^{7+} , Ne
- 6) Isobaric → Isobaric species having same no. of atoms/molecules.



Thomson's Model of Atom :-

- Also known as plum pudding / watermelon model.
- According to J.J. Thomson, atom consists of uniform sphere of +ve charge in which the electrons are distributed uniformly.



Drawbacks :-

- Mass considered evenly spread over the atom.
- It is a stable model. Doesn't explain movement of e^- .
- Can't explain stability of atom.
- Failed to explain Rutherford α -particle experiment.

Rutherford Model & α -particle Experiment :-

Rutherford bombarded α -particle on a thin gold foil.



Observations :-

- Most of the α -particles pass through the gold foil undeviated.
- Few α -particles get deflected by some angle.

→ Very few α -particles (1 in 2000) returned back to its original path.

Conclusions :-

- 1) Most of the atom is empty.
- 2) Major part by mass concentrated at center in very small region called 'nucleus'.
- 3) Nucleus is very small as compared to atom.

$$r_{\text{atom}} \approx 10^{-10} \text{ m}; \quad r_{\text{nucleus}} \approx 10^{-15} \text{ m}$$

- 4) Nucleus is +vely charged.
- 5) Protons and neutrons are present in nucleus.
- 6) Since an atom is neutral, no. of proton = no. of e^- .
- 7) e^- revolves around the nucleus in a definite path called orbits.

Distance of Closest Approach :-

→ Minimum distance from nucleus at which α -particle stops due to repulsion & goes back.



$$R = \frac{KZe^2}{mv^2}$$

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