

5-2 Review and Reinforcement Part II

Quantum Theory and the Atom

Complete each statement, using the vocabulary and the characteristics and history of quantum theory.

characteristic

Pauli exclusion principle

quantum-mechanical model

quantum-mechanical

electron density

orbital

principal energy levels

_____ orbital _____ 1. A region in space around the nucleus where electrons in atoms are likely to be.

_____ electron density _____ 2. The density of an electron cloud.

_____ quantum-mechanical _____ 3. Another description of a quantum-mechanical model of an atom.

Pauli exclusion principle is a rule that says that no two electrons in an atom can have the same _____ and the same _____.

quantum-mechanical model; spin

_____ quantum-mechanical _____ 4. Another description of an atom, by describing the electron density and _____.

quantum-mechanical model

_____ principal energy levels _____ 5. The levels of energy that electrons _____.

_____ electron spin _____ 6. The clockwise or counterclockwise rotation of an electron.

Complete each statement, using the history of the quantum-mechanical model of the atom.

_____ 1. _____ The electron cloud is _____ where the probability of finding an electron is _____.

- a. greatest
- b. lowest
- c. highest density
- d. intermediate

_____ 2. _____ The low principal energy level of the hydrogen atom contains only one _____.

- a. orbital
- b. s orbital
- c. d orbital
- d. p orbital

20 2 Review And Reinforcement Answers Chemistry

**Robert C. Smoot, Smoot, Richard G.
Smith, Jack Price**



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