

Acids and Bases Worksheet

1. Acids and bases share some properties but are different in others. Use the samples at the side of the class to complete the chart. If both acids and bases share the property, place it in both columns.

Property	Acids	Bases
Conductivity in Aqueous Solution	very	very
Taste	Sour	Bitter
Feel	—	Slippery
Reaction with Metal	produces H_2 gas	—
Reaction with Carbonate	produces CO_2 gas	—
Colour on Litmus Paper	Red	Blue
Colour with Universal Indicator	Red	Purple
Colour with phenolphthalein	Colourless	pink

2. Indicate if each compound is an acid or base and name it.

	Acid/Base?	Name
HNO_3	<u>A</u>	<u>nitric acid</u>
$Mg(OH)_2$	<u>B</u>	<u>magnesium hydroxide</u>
H_3PO_4	<u>A</u>	<u>phosphoric acid</u>
H_2CO_3	<u>A</u>	<u>carbonic acid</u>
HF	<u>A</u>	<u>hydrofluoric acid</u>
$Fe(OH)_3$	<u>B</u>	<u>iron (iii) hydroxide</u>

3. Write the formula for each of the following acids/bases:

hydrochloric acid	<u>HCl</u>	hydrochloric acid	<u>HCl</u>
aluminum hydroxide	<u>$Al(OH)_3$</u>	sodium hydroxide	<u>NaOH</u>
lithium hydroxide	<u>LiOH</u>	calcium hydroxide	<u>$Ca(OH)_2$</u>
sulfuric acid	<u>H_2SO_4</u>	sulfurous acid	<u>H_2SO_3</u>

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Table of Contents 20 Describing Acids And Bases Answer Key

1. Understanding the eBook 20 Describing Acids And Bases Answer Key
 - The Rise of Digital Reading 20 Describing Acids And Bases Answer Key
 - Advantages of eBooks Over Traditional Books
2. Identifying 20 Describing Acids And Bases Answer Key
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an 20 Describing Acids And Bases Answer Key
 - User-Friendly Interface
4. Exploring eBook Recommendations from 20 Describing Acids And Bases Answer Key
 - Personalized Recommendations
 - 20 Describing Acids And Bases Answer Key User Reviews and Ratings
 - 20 Describing Acids And Bases Answer Key and Bestseller Lists
5. Accessing 20 Describing Acids And Bases Answer Key Free and Paid eBooks

- 20 Describing Acids And Bases Answer Key Public Domain eBooks
 - 20 Describing Acids And Bases Answer Key eBook Subscription Services
 - 20 Describing Acids And Bases Answer Key Budget-Friendly Options
6. Navigating 20 Describing Acids And Bases Answer Key eBook Formats
- ePub, PDF, MOBI, and More
 - 20 Describing Acids And Bases Answer Key Compatibility with Devices
 - 20 Describing Acids And Bases Answer Key Enhanced eBook Features
7. Enhancing Your Reading Experience
- Adjustable Fonts and Text Sizes of 20 Describing Acids And Bases Answer Key
 - Highlighting and Note-Taking 20 Describing Acids And Bases Answer Key
 - Interactive Elements 20 Describing Acids And Bases Answer Key
8. Staying Engaged with 20 Describing Acids And Bases Answer Key
- Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers 20 Describing Acids And Bases Answer Key
9. Balancing eBooks and Physical Books 20 Describing Acids And Bases Answer Key
- Benefits of a Digital Library
 - Creating a Diverse Reading Collection 20 Describing Acids And Bases Answer Key
10. Overcoming Reading Challenges
- Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine 20 Describing Acids And Bases Answer Key
- Setting Reading Goals 20 Describing Acids And Bases Answer Key
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of 20 Describing Acids And Bases Answer Key
- Fact-Checking eBook Content of 20 Describing Acids And Bases Answer Key
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
- Utilizing eBooks for Skill Development

- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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