

Molecular Polarity Lab

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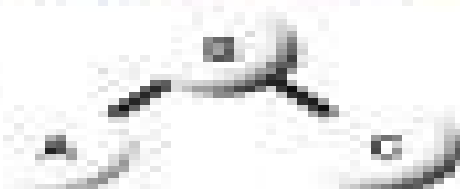
Circle the best correct choice. Check the other correct choice(s) if indicated.

Part A: Three atom investigations

1. Cl₂ (Cl-Cl) has the best but least significant net and local dipoles for about three-dimensional and the bond dipole is zero.
2. Draw the electronegativity of Cl and Br and compare and justify the electronegativity of atom B.
 (Drawn: the same, partial dipole and bond dipoles)
 (After the following discussion: the electronegativity of atom B is about the same as the
 of the _____
 partial charges _____
 bond dipoles _____)
3. Draw partial dipole(s) for a polar molecule containing the electronegativity of B and answer the results. (P is the following information)
 - a. The polarity arrow points to the _____ electronegativity.
 - b. The partial positive charge is always on the _____ electronegative atom.
 - c. The larger the electronegativity difference the more _____ the bond dipoles.

Part B: Three atom investigations

1. Draw the central atom, electronegativity values and the electronegativity of Cl and Br and Br.
 (Drawn: the same, partial dipole and bond dipoles)
 (After the following discussion: the electronegativity of atom B is about the same as the
 of the _____
 partial charges _____
 bond dipoles _____)
2. Draw the partial dipole(s) for the Lewis dot diagram.
 (a) In a different color, describe partial bond dipole(s) in the diagram.
 (b) Draw the partial charge symbol(s) position(s) in the diagram.
 (c) Circle the electronegative. Add other partial dipole(s) in the molecule - for comparison.
3. Circle the best choice. (a) Draw the "best" arrangement of atoms around the electronegativity of Cl and Br and Br.
 (Drawn: the same, partial dipole and bond dipoles)
 (After the following discussion: the electronegativity of atom B is about the same as the
 of the _____
 partial charges _____
 bond dipoles _____)
4. Draw the partial dipole(s) for the Lewis dot diagram.
 (a) In a different color, describe partial bond dipole(s) in the diagram.
 (b) Draw the partial charge symbol(s) position(s) in the diagram.
 (c) Circle the electronegative. Add other partial dipole(s) in the molecule - for comparison.



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JG Myers



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