

# Hydrolysis -

## The Reaction of Ions with Water

### Lab Report

#### Purpose

The purpose of this lab is to predict whether a substance is acidic, basic, or neutral and to use the universal indicator solution to test it.

#### Hypothesis

|                   |                                                                                                                                                                                                                                                                                                                                                                                                                                                         |                             |         |
|-------------------|---------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------|-----------------------------|---------|
| Sodium Acetate    | $\text{NaCH}_3\text{COO}_{(s)} \rightarrow \text{Na}^+_{(aq)} + \text{CH}_3\text{COO}^-_{(aq)}$ $\text{CH}_3\text{COO}^-_{(aq)} + \text{H}(\text{OH})_{(l)} \leftrightarrow \text{CH}_3\text{COOH}_{(aq)} + \text{OH}^-_{(aq)}$                                                                                                                                                                                                                         |                             | Basic   |
| Sodium Chloride   | $\text{NaCl}_{(s)} \leftrightarrow \text{Na}^+_{(aq)} + \text{Cl}^-_{(aq)}$                                                                                                                                                                                                                                                                                                                                                                             |                             | Neutral |
| Ammonium Chloride | $\text{NH}_4\text{Cl}_{(s)} \rightarrow \text{NH}_4^+_{(aq)} + \text{Cl}^-_{(aq)}$ $\text{NH}_4^+_{(aq)} + \text{H}(\text{OH})_{(l)} \leftrightarrow \text{NH}_4(\text{OH})_{(aq)} + \text{H}^+_{(aq)}$                                                                                                                                                                                                                                                 |                             | Acidic  |
| Ammonium Sulfate  | $(\text{NH}_4)_2\text{SO}_{4(s)} \rightarrow \text{NH}_4^+_{(aq)} + \text{SO}_4^{2-}_{(aq)}$ $\text{NH}_4^+_{(aq)} + \text{H}(\text{OH})_{(l)} \leftrightarrow \text{NH}_4(\text{OH})_{(aq)} + \text{H}^+_{(aq)}$ $\text{SO}_4^{2-}_{(aq)} + \text{H}(\text{OH})_{(l)} \leftrightarrow \text{HSO}_4^-_{(aq)} + \text{OH}^-_{(aq)}$                                                                                                                      | $K_a = 5.4 \times 10^{-10}$ | Acidic  |
|                   |                                                                                                                                                                                                                                                                                                                                                                                                                                                         | $K_b = 8.3 \times 10^{-13}$ |         |
| Aluminum Chloride | $\text{AlCl}_{3(s)} \rightarrow \text{Al}^{3+}_{(aq)} + \text{Cl}^-_{(aq)}$ $\text{Al}^{3+}_{(aq)} + \text{H}(\text{OH})_{(l)} \leftrightarrow \text{Al}(\text{OH})_{(aq)} + \text{H}^+_{(aq)}$                                                                                                                                                                                                                                                         |                             | Acidic  |
| Calcium Nitrate   | $\text{Ca}(\text{NO}_3)_{2(s)} \rightarrow \text{Ca}^{2+}_{(aq)} + \text{NO}_{3(aq)}$                                                                                                                                                                                                                                                                                                                                                                   |                             | Neutral |
| Iron(III) Sulfate | $\text{Fe}_2(\text{SO}_4)_{3(s)} \rightarrow \text{Fe}^{3+}_{(aq)} + \text{SO}_4^{2-}_{(aq)}$ $\text{Fe}^{3+}_{(aq)} + \text{H}(\text{OH})_{(l)} \leftrightarrow \text{Fe}(\text{OH})^{2+}_{(aq)} + \text{H}^+_{(aq)}$ $\text{Fe}^{3+}_{(aq)} + \text{H}(\text{OH})_{(l)} \leftrightarrow \text{HFe}^{2+}_{(aq)} + \text{OH}^-_{(aq)}$ $\text{SO}_4^{2-}_{(aq)} + \text{H}(\text{OH})_{(l)} \leftrightarrow \text{HSO}_4^-_{(aq)} + \text{OH}^-_{(aq)}$ | $K_a = 6.0 \times 10^{-3}$  | Acidic  |
|                   |                                                                                                                                                                                                                                                                                                                                                                                                                                                         | $K_b = 1.7 \times 10^{-12}$ |         |
|                   |                                                                                                                                                                                                                                                                                                                                                                                                                                                         | $K_b = 8.3 \times 10^{-13}$ |         |
| Sodium Carbonate  | $\text{Na}_2\text{CO}_{3(s)} \rightarrow \text{Na}^+_{(aq)} + \text{CO}_3^{2-}_{(aq)}$ $\text{CO}_3^{2-}_{(aq)} + \text{H}(\text{OH})_{(l)} \leftrightarrow \text{HCO}_3^-_{(aq)} + \text{OH}^-_{(aq)}$                                                                                                                                                                                                                                                 |                             | Basic   |

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## **Table of Contents Answer Of Heath Chemistry Lab 20**

1. Understanding the eBook Answer Of Heath Chemistry Lab 20
  - The Rise of Digital Reading Answer Of Heath Chemistry Lab 20
  - Advantages of eBooks Over Traditional Books
2. Identifying Answer Of Heath Chemistry Lab 20
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Answer Of Heath Chemistry Lab 20
  - User-Friendly Interface
4. Exploring eBook Recommendations from Answer Of Heath Chemistry Lab 20
  - Personalized Recommendations
  - Answer Of Heath Chemistry Lab 20 User Reviews and Ratings
  - Answer Of Heath Chemistry Lab 20 and Bestseller Lists
5. Accessing Answer Of Heath Chemistry Lab 20 Free and Paid eBooks

- Answer Of Heath Chemistry Lab 20 Public Domain eBooks
- Answer Of Heath Chemistry Lab 20 eBook Subscription Services
- Answer Of Heath Chemistry Lab 20 Budget-Friendly Options
- 6. Navigating Answer Of Heath Chemistry Lab 20 eBook Formats
  - ePub, PDF, MOBI, and More
  - Answer Of Heath Chemistry Lab 20 Compatibility with Devices
  - Answer Of Heath Chemistry Lab 20 Enhanced eBook Features
- 7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Answer Of Heath Chemistry Lab 20
  - Highlighting and Note-Taking Answer Of Heath Chemistry Lab 20
  - Interactive Elements Answer Of Heath Chemistry Lab 20
- 8. Staying Engaged with Answer Of Heath Chemistry Lab 20
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Answer Of Heath Chemistry Lab 20
- 9. Balancing eBooks and Physical Books Answer Of Heath Chemistry Lab 20
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Answer Of Heath Chemistry Lab 20
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine Answer Of Heath Chemistry Lab 20
  - Setting Reading Goals Answer Of Heath Chemistry Lab 20
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Answer Of Heath Chemistry Lab 20
  - Fact-Checking eBook Content of Answer Of Heath Chemistry Lab 20
  - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development

- Exploring Educational eBooks

#### 14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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