



Equations

Name _____

Date _____

Write a balanced chemical equation to represent each of the following chemical reactions.

1. iron + sulfur $\rightarrow$ iron(II) sulfide
2. zinc + copper(II) sulfate $\rightarrow$ zinc sulfate + copper
3. silver nitrate + sodium bromide $\rightarrow$ sodium nitrate + silver bromide
4. potassium chlorate (heated) $\rightarrow$ potassium chloride + oxygen
5. water (electricity) $\rightarrow$ hydrogen + oxygen
6. mercury(II) oxide (heated) $\rightarrow$ mercury + oxygen
7. potassium iodide + lead(II) nitrate $\rightarrow$ lead(II) iodide + potassium nitrate
8. aluminum + oxygen $\rightarrow$ aluminum oxide
9. magnesium chloride + ammonium nitrate $\rightarrow$ magnesium nitrate + ammonium chloride
10. iron(III) chloride + ammonium hydroxide $\rightarrow$ iron(III) hydroxide + ammonium chloride
11. sodium peroxide + water $\rightarrow$ sodium hydroxide + oxygen
12. iron(III) oxide + carbon $\rightarrow$ iron + carbon monoxide
13. iron + water $\rightarrow$ hydrogen + iron(III) oxide
14. iron(III) chloride + potassium hydroxide $\rightarrow$ potassium chloride + iron(III) hydroxide
15. aluminum + sulfuric acid $\rightarrow$ aluminum sulfate + hydrogen
16. sodium carbonate + calcium hydroxide $\rightarrow$ sodium hydroxide + calcium carbonate



Answer Key Weston Walch Chemistry

Judd E. Hollander



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