

অধ্যায় ১: ল্যাবরেটরির নিরাপদ ব্যবহার



পটীক্ষার কপূর, এখানে সেরে পড়িষ্ঠা, আরও কলেজের শিক্ষার্থীর কলেজসমূহের পটীক্ষা এবং মতামত টিউটর পটীক্ষার পূর্বের উত্তর অধ্যয়নিক সেবা করে। এগুলোর অনুশীলন করলে তুমি এ অধ্যয়ন থেকে সেরে পড়িষ্ঠা পূর্ণাঙ্গীত করবে।

প্রশ্ন ১



পাত্র-A



পাত্র-B

সি. কে. ১০/১০/১০

- মোলের প্রকণ কী?
- "কেন্দ্র পানিযুক্ত লবণের প্রকণতা, 'আপনাত্মা বৃষ্টির সাথে সাথে এক পর্যায়ে বৃষ্টি না পেয়ে হ্রাস পায়'— ব্যাখ্যা করে।
- A-পাত্রের বিন্যাসন এলিটের ভর কত নির্ণয় করে।
- প্রকণ দুটি মিশ্রিত করলে, মিশ্রণটি কী পরিমাণ অক্সিজেন ফাঁকীয় হবে— বিশ্লেষণ করে।

১ নং প্রশ্নের উত্তর

1000 mL প্রকণে কোন প্রকণের 1 mol প্রকণিত থাকলে উক্ত প্রকণকে মোলার প্রকণ বলে।

কেন্দ্র পানিযুক্ত লবণের প্রকণতা আপনাত্মা বৃষ্টির সাথে সাথে এক পর্যায়ে বৃষ্টি না পেয়ে হ্রাস পায়। এর কারণ, আপনাত্মা বৃষ্টির সাথে সাথে প্রকণতা বৃষ্টি পেতে থাকলেও একটি নির্দিষ্ট আপনাত্মার কেন্দ্র পানি দুই লবণ নিষ্কৃতি হয়। তখন প্রকণতা হ্রাস পেতে থাকে। উদাহরণস্বরূপ $\text{Na}_2\text{SO}_4 \cdot 10\text{H}_2\text{O}$ এর কথা বলা যায়। $\text{Na}_2\text{SO}_4 \cdot 10\text{H}_2\text{O}$ এর কেন্দ্র পানিতে প্রকণিত করার সময় বিশ্লেষণ প্রক্রিয়াটি আপনাত্মার Na_2SO_4 এ পরিণত হয় তখন বিশ্লেষণ প্রক্রিয়াটি আপোৎপাদী হয়। এ কারণে 32.4°C আপনাত্মার উপরে আপনাত্মা বৃষ্টিজনিত কারণে প্রকণতার হ্রাস ঘটে।

১

আমরা জানি,
 $w = SMV$
 $= (0.5 \times 98 \times 0.25) \text{ g}$
 $= 12.25 \text{ g}$
 অতএব, প্রকণিত H_2SO_4 এর ভর
 $= 12.25 \text{ g}$

উদীপকের A-পাত্রের অবস্থিত,
 H_2SO_4 এর আণবিক ভর,
 $M = 1 \times 2 + 32 + 16 \times 4$
 $= 2 + 32 + 64$
 $= 98 \text{ g/mol}$
 প্রকণের ঘনমাত্রা,
 $S = 0.5 \text{ M}$
 প্রকণের আয়তন,
 $V = 250 \text{ mL}$
 $= \frac{250}{1000} = 0.25 \text{ L}$

প্রকণিত H_2SO_4 এর ভর, $w = ?$

প্রকণ-A: H_2SO_4 এর মোল সংখ্যা $= \left(0.5 \times \frac{250}{1000}\right) \text{ mol}$
 $= 0.125 \text{ mol}$

প্রকণ-B: Na_2CO_3 এর মোল সংখ্যা $= \left(0.11 \times \frac{250}{1000}\right) \text{ mol}$
 $= 0.077 \text{ mol}$

সম্প্রতি বিক্রিয়া: $\text{H}_2\text{SO}_4 + \text{Na}_2\text{CO}_3 \rightarrow \text{Na}_2\text{SO}_4 + \text{H}_2\text{O} + \text{CO}_2$

সমীকরণ হতে পাই, 1 mol $\text{H}_2\text{SO}_4 = 1 \text{ mol Na}_2\text{CO}_3$

1 mol Na_2CO_3 বিক্রিয়া করে H_2SO_4 এর 1 mol এর সাথে
 $0.077 \text{ mol} = 0.077 \text{ mol} = 0.077 \text{ mol} = 0.077 \text{ mol}$

অবশিষ্ট H_2SO_4 এর পরিমাণ $= (0.125 - 0.077) \text{ mol}$
 $= 0.048 \text{ mol}$

পাত্র-A এর প্রকণের সাথে পাত্র-B এর প্রকণের বিক্রিয়ার ফলে 0.048 mol H_2SO_4 অবশিষ্ট থাকায় মিশ্রিত প্রকণের প্রকৃতি হবে অক্সিজেন।

প্রশ্ন ২ এককাল ছাত্র ল্যাবরেটরিতে X-পদ্ধতিতে 10 mg এর প্রকণ পরিমাণ এবং Y-পদ্ধতিতে 0.5 g পরিমাণ 2% $\frac{W}{V}$ HCl এর নমুনা কাল করে।

সি. কে. ১০/১০/১০

- নোড কাকে বলে?
- Cl_2 অপোনার— ব্যাখ্যা করে।
- উদীপকের এলিটটির মোলার ঘনমাত্রা নির্ণয় করে।
- উদীপকের কোন পদ্ধতি পরিবেশবান্ধব? বিশ্লেষণ করে।

২ নং প্রশ্নের উত্তর

নিউক্লিয়াসের চারদিকে যে নির্দিষ্ট ত্রিমাত্রিক স্থানে কোন ইলেক্ট্রন পাওয়ার সম্ভাব্যতা শূন্য তাকে নোড বলে।

সমযোজী বন্ধনে আবদ্ধ উভয় পরমাণুর তড়িৎ ঋণাত্মকতা হল অর্ধাং তড়িৎ ঋণাত্মকতার পার্থক্য শূন্য হলে সমযোজী বন্ধন শূন্য ও অপোনার হয়। আর তড়িৎ ঋণাত্মকতার পার্থক্য 0.5 এর কম হলে অপোনার হয়। যেহেতু Cl এর তড়িৎ ঋণাত্মকতা 3.0 ও C অণুতে তড়িৎ ঋণাত্মকতার পার্থক্য $(3 - 3) = 0$, তাই Cl_2 বিশুদ্ধ সমযোজী ও অপোনার যৌগ। উল্লেখ্য যৌগিক অণুর ক্ষেত্রে এহুণ থাকে।

উদীপকে বর্ণিত এলিটটি হলো 2% $\left(\frac{W}{V}\right) \text{ HCl}$

2% $\left(\frac{W}{V}\right) \text{ HCl}$ এর অর্থ হলো—

100 mL প্রকণে HCl এর ভর 2g

∴ 1000 mL " " " " $\left(\frac{2 \times 1000}{100}\right) \text{ g}$
 $= 20 \text{ g}$

Answer Of Hsc Chemistry 1st Paper 15 Ctg Board

Jin-Ying Zhang



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