

## QUESTION 2

|                   | Test                                                                               | Observation                                                                                                         | Inference                                       |
|-------------------|------------------------------------------------------------------------------------|---------------------------------------------------------------------------------------------------------------------|-------------------------------------------------|
|                   | (a)(i) C + Conc. $\text{H}_2\text{SO}_4$ + strong heat                             | A copious brown irritating fume which turned moist blue litmus paper to red                                         | Gas is $\text{NO}_2$<br>$\text{NO}_3^-$ present |
|                   | (b)(i) C + water + NaOH in drops, in excess                                        | white gelatinous precipitate<br>precipitate dissolved                                                               | $\text{Al}^{3+}$ or $\text{Zn}^{2+}$            |
|                   | (ii) C + water + $\text{NH}_3$ in drops, in excess                                 | white gelatinous precipitate<br>precipitate dissolved                                                               | $\text{Al}^{3+}$ or $\text{Zn}^{2+}$            |
| Trickle down test | (iii) C + water + freshly prepared $\text{FeSO}_4$ + Conc. $\text{H}_2\text{SO}_4$ | A brown ring is formed at the junction of the acid and aqueous solution.                                            | $\text{NO}_3^-$                                 |
|                   | (c)(i) D + water                                                                   | D dissolved                                                                                                         | soluble                                         |
|                   | (ii) Solution D + dil HCl                                                          | Colourless gas, smells like rotten egg, turns moist blue litmus paper to red and lead (II) ethanoate paper to black | $\text{H}_2\text{S}$ present                    |
|                   | (iii) Solution D + NaOH in drops, in excess                                        | white precipitate<br>precipitate insoluble                                                                          | $\text{Ca}^{2+}$ or $\text{Pb}^{2+}$            |
|                   | (iv) solution D + $(\text{NH}_4)_2\text{CO}_3$                                     | white precipitate                                                                                                   | $\text{Ca}^{2+}$                                |

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