

**AP[®] CHEMISTRY
2019 SCORING GUIDELINES**

Question 1



Ethanol, $\text{C}_2\text{H}_5\text{OH}$, will combust in air according to the equation above.

(a) Is $\text{O}_2(g)$ oxidized in the reaction, or is it reduced? Justify your answer in terms of oxidation numbers.

O_2 is reduced; its oxidation number changes from 0 in O_2 to -2 in CO_2 and H_2O .

1 point is earned for the correct answer with an appropriate justification.

(b) When a sample of $\text{C}_2\text{H}_5\text{OH}$ was combusted, the volume of $\text{CO}_2(g)$ produced was 18.0 L when measured at 21.7°C and 1.03 atm. Determine the number of moles of $\text{CO}_2(g)$ that was produced.

$$n = \frac{PV}{RT} = \frac{(1.03 \text{ atm})(18.0 \text{ L})}{(0.08206 \text{ L atm mol}^{-1} \text{ K}^{-1})(273.15 + 21.7 \text{ K})} = 0.766 \text{ mol CO}_2$$

1 point is earned for the correct calculation.

(c) Determine the volume of $\text{C}_2\text{H}_5\text{OH}(l)$, in mL, that was combusted to produce the volume of $\text{CO}_2(g)$ collected in part (b). (The density of $\text{C}_2\text{H}_5\text{OH}(l)$ is 0.79 g/mL.)

$$0.766 \text{ mol CO}_2 \times \frac{1 \text{ mol C}_2\text{H}_5\text{OH}}{2 \text{ mol CO}_2} \times \frac{46.07 \text{ g C}_2\text{H}_5\text{OH}}{1 \text{ mol C}_2\text{H}_5\text{OH}} \times \frac{1 \text{ mL}}{0.79 \text{ g}} = 22 \text{ mL}$$

1 point is earned for the correct stoichiometry for moles of $\text{C}_2\text{H}_5\text{OH}$ and CO_2 .

1 point is earned for the correct volume of $\text{C}_2\text{H}_5\text{OH}$.

(d) Determine the amount of heat, in kJ, that is released by the combustion reaction.

$$q = n \Delta H_{\text{comb}}^\circ$$

$$q = 0.766 \text{ mol CO}_2 \times \frac{1 \text{ mol}_{\text{C}_2\text{H}_5\text{OH}}}{2 \text{ mol CO}_2} \times \frac{-1270 \text{ kJ}}{1 \text{ mol}_{\text{C}_2\text{H}_5\text{OH}}} = -486 \text{ kJ}$$

The amount of heat released is 486 kJ.

1 point is earned for a correct calculation (negative sign is not required).

The combustion reaction occurred in a closed room containing $5.56 \times 10^4 \text{ g}$ of air originally at 21.7°C. Assume that all of the heat produced by the reaction was absorbed by the air (specific heat = 1.005 J/g °C) in the room.

20ap Chem Frq Scoring Guidelines

John T. Moore, Richard H. Langley



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Table of Contents 20ap Chem Frq Scoring Guidelines

1. Understanding the eBook 20ap Chem Frq Scoring Guidelines
 - The Rise of Digital Reading 20ap Chem Frq Scoring Guidelines
 - Advantages of eBooks Over Traditional Books
2. Identifying 20ap Chem Frq Scoring Guidelines
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an eBook Platform
 - User-Friendly Interface
4. Exploring eBook Recommendations from 20ap Chem Frq Scoring Guidelines
 - Personalized Recommendations
 - 20ap Chem Frq Scoring Guidelines User Reviews and Ratings
 - 20ap Chem Frq Scoring Guidelines and Bestseller Lists
5. Accessing 20ap Chem Frq Scoring Guidelines Free and Paid eBooks
 - 20ap Chem Frq Scoring Guidelines Public Domain eBooks
 - 20ap Chem Frq Scoring Guidelines eBook Subscription Services
 - 20ap Chem Frq Scoring Guidelines Budget-Friendly Options
6. Navigating 20ap Chem Frq Scoring Guidelines eBook Formats
 - ePub, PDF, MOBI, and More
 - 20ap Chem Frq Scoring Guidelines Compatibility with Devices
 - 20ap Chem Frq Scoring Guidelines Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of 20ap Chem Frq Scoring Guidelines
 - Highlighting and Note-Taking 20ap Chem Frq Scoring Guidelines
 - Interactive Elements 20ap Chem Frq Scoring Guidelines
8. Staying Engaged with 20ap Chem Frq Scoring Guidelines

- Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers 20ap Chem Frq Scoring Guidelines
9. Balancing eBooks and Physical Books 20ap Chem Frq Scoring Guidelines
- Benefits of a Digital Library
 - Creating a Diverse Reading Collection 20ap Chem Frq Scoring Guidelines
10. Overcoming Reading Challenges
- Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine 20ap Chem Frq Scoring Guidelines
- Setting Reading Goals 20ap Chem Frq Scoring Guidelines
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of 20ap Chem Frq Scoring Guidelines
- Fact-Checking eBook Content of 20ap Chem Frq Scoring Guidelines
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
- Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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