পরীকারী কপুরা, এখানে বোর্র পরীকা, ক্যাডোট কলেলগর শীর্ষান্ধানীর কলেলগমূহের পরীক্ষা এবং মহেল টেন্টের ব্যাত্তর পূর্ণাল উত্তর অধ্যায়ভিত্তিক দেবছা ময়েছে। এগুলো অনুশীলন করনে কৃষি এ অধ্যায়টি লেকে সেকোনো স্থানশীল জনামূদ্





ক, মোলার দ্রবণ কীৰু

 "কেলাস পানিযুক্ত লকদের দ্রাক্তার, তাপমাত্রা বুল্পির সাথে সাথে এক পর্যায়ে বৃশ্বি না পেয়ে হ্রাস পায়"— ব্যাখ্যা করে।

ণ. A-পাত্রে বিদর্শান এগিজের ভর কত নির্ণয় করো।

 ছবল দুটি মিশ্রিত করলে, থিপ্রদটি কী পরিমাণ অগ্নীয় বা ফারীয় মবে— বিশ্লেষণ করো।

১ নং প্রয়ের উত্তর

1000 ml. চবণে কোন স্তবের 1 mol স্তবীভূত থাকলে উদ্ভ স্তবগকে
মোলার স্তবণ কলে।

ত্রনাস পানিত্র গবণের প্রাব্যান্তা তাপমাত্রা বৃশ্বির সাথে সাথে এক পর্যান্ত বৃশ্বির না পেরা ছাস পার। এর কারণ, তাপমাত্রা বৃশ্বির সাথে সাথে রাক বারণ, তাপমাত্রা বৃশ্বির সাথে সাথে প্রকাশ রাক বারণ রাক পেতে থাকলের একটি নির্নিন্দ তাপমাত্রার কেলাস পানি কুর গবণ নির্নান্ত হয়। তথন প্রাব্যান্তা স্ত্রাস পেতে থাকে। তদাহরণমূপ Na₂SO₄.10H₂O এর কথা করা যায়। Na₂SO₄.10H₂O এর পেতে পানিতে প্রবীত্ত হওয়ার সময় বিয়োজন প্রক্রিয়ান্তি ভাপহারী বিক্রিয়া কিন্তু 32.4°C তাপমাত্রার উপরে Na₂SO₄.10H₂O যখন Na₂SO₄ পরিণত হয় তথন বিয়োজন প্রক্রিয়াটি ভাপোৎপাদী হয়। এ কারণে 32.4°C তাপমাত্রার উপরে তাপমাত্রা কৃশ্বিজনিত কারণে প্রাব্যান্ত স্থান হল হল হল হটে।

कामदा कानि

w = SMV = (0.5 × 98 × 0.25) g = 12.25 g অভন্তন্ত্র, দ্রবীভূত H₂SO₄ এর ভর

উদ্দীপকের A-পাত্রে অবস্থিত,

H₂SO₄ এর আণবিক ভর,

M = 1 × 2 + 32 + 16 × 4

= 2 + 32 + 64

= 98 g/mol

রবণের ঘনমারা,

S = 0.5 M

রবণের আয়তন,

V = 250 mL

= 250

1000 = 0.25L

ফবীভূত H₂SO₄ এর ভর, w = 1

ভাল-A: H₂SO₄ গান্ত দোলা সংখ্যা — $\left(0.5 \times \frac{250}{1000}\right)$ med — 0.125 mod — 0.125 mod — 0.125 mod — 0.077 mod — 0.077 mod — 0.077 mod সংঘটিত বিভিন্না: H₂SO₄ + Na₂CO₃ → Na₂SO₄ + H₂O + CO₃ স্থীকরণ হতে পাই, I mod H₂SO₄ = I mod Na₃CO₃ I mod Na₂CO₃ বিভিন্না করে H₂SO₄ এই I mod এই সংস্থা — $\left(0.077\right)$ mod — — $\left(0.077$

19.05

ett.

পাত্র–∧ এর স্তবণের সাথে পাত্র–৪ এর প্রবণের বিভিন্নার ছলে ১৯ mol II₂SO, অবশিউ থাকায় মিশ্রিত স্তবণের প্রকৃতি ছবে অভীত।

ভর্মা > ব্যক্তন ছাত্র লাবরেটরিতে X-পশ্বতিতে 10 mg বর তেওঁ হ পরিমাণ এবং Y-পশ্বতিতে 0.5 g পরিমাণ 2% ₩ HCl এর নমূর হ কাজ করে।

ক. লোভ কাকে বলে?

খ. Cl₂ অপোলার- ব্যাখ্যা করে।

ণ, উন্দীপকের এসিভটির মোলার ঘনমাতা নির্ণয় করে।

য়, উদ্দীপকের কোন পশ্বতি পরিবেশবান্ধর? বিশ্লেষণ করে। ২ নং প্রয়ের উত্তর

নিউরিয়াসের চারলিকে যে নির্দিট ত্রিমাত্রিক স্থানে কানে ইজক পাওয়ার সম্ভাব্যতা শূন্য তাকে নোভ বলে।

সমযোজী বন্ধনে আবন্ধ উভয় পরমাণুর তড়িং কণাছকত ক হলে অর্থাৎ তড়িং ঝণাঝকতার পার্থকা শূন্য হলে সমযোজী ক বিশুন্ধ ও অপোলার হয়। আর তড়িং ঝণাঝকতার পার্থকা 0.5 এর ক মর্দে অপোলার হয়। মেহেতু CI এর তড়িং ঝণাঝকতা 3.0 ৫ ট অপুতে তড়িং ঝণাঝকতার পার্থকা (3 – 3)= 0, তাই CI₂ বিশ্ব সমযোজী ও অপোলার যৌগ। উল্লেখ্য মৌলিক অপুর ক্ষেত্রে এমুশ হা থাকে।

ত্রী উদ্দীপকে বর্ণিত এগিচটি হলো 2% $\left(\frac{W}{V}\right)$ HCI ।

2% $\left(\frac{W}{V}\right)$ HCI এর অর্থ ফলো- 100 mL দ্রবংশ HCI এর ভর 2ϵ

.. 1000 mL " " (2 × 1000) 100 g

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Wenbin Ji

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