

Which of the following mixtures is a buffer with a pH of approximately 7?

- (A) A mixture of 100. mL of 0.1 M  $\text{CH}_3\text{NH}_3\text{Cl}$  and 50. mL of 0.1 M NaOH
- (B) A mixture of 100. mL of 0.1 M  $\text{HC}_2\text{H}_3\text{O}_2$  and 50. mL of 0.1 M NaOH
- (C) A mixture of 100. mL of 0.1 M  $\text{NaC}_2\text{H}_3\text{O}_2$  and 100. mL of 0.1 M NaOH
- (D) A mixture of 100. mL of 0.1 M  $\text{CH}_3\text{NH}_3\text{Cl}$  and 100. mL of 0.1 M  $\text{CH}_3\text{NH}_2$



36. The Lewis electron-dot diagrams of the  $\text{HClO}_3$  molecule and the  $\text{HClO}_4$  molecule are shown above at the left and right, respectively. Which of the following statements identifies the stronger acid and correctly identifies a factor that contributes to its being the stronger acid?

- (A)  $\text{HClO}_3(\text{aq})$  is the stronger acid because its molecules experience stronger London dispersion forces.
- (B)  $\text{HClO}_3(\text{aq})$  is the stronger acid because the additional electronegative oxygen atom on the chlorine atom stabilizes the conjugate base.
- (C)  $\text{HClO}_4(\text{aq})$  is the stronger acid because its



40. The reaction system represented above is at equilibrium. Which of the following will decrease the amount of  $\text{CaOH}(\text{s})$  in the system?

- (A) Increasing the volume of the reaction vessel at constant temperature
- (B) Lowering the temperature of the system
- (C) Removing some  $\text{CO}_2(\text{g})$  at constant temperature
- (D) Removing some  $\text{CaCO}_3(\text{s})$  at constant temperature

| Acid Solution | Volume of NaOH Added (mL) |
|---------------|---------------------------|
| A             | 40                        |
| B             | 75                        |
| C             | 115                       |
| D             | 200                       |

$$M_A V_A = M_B V_B$$

41. To maximize the yield in a certain manufacturing process, a solution of a weak monoprotic acid that has a concentration between 0.20 M and 0.30 M is required. Four 100. mL samples of the acid at different concentrations are each titrated with a 0.20 M NaOH solution. The volume of NaOH needed to reach the end point for each sample is shown in the table above. Which solution is the

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**J Ma**

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## **Table of Contents Ap Chemistry Multiple Choice Explanation**

1. Understanding the eBook Ap Chemistry Multiple Choice Explanation
  - The Rise of Digital Reading Ap Chemistry Multiple Choice Explanation
  - Advantages of eBooks Over Traditional Books
2. Identifying Ap Chemistry Multiple Choice Explanation
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Ap Chemistry Multiple Choice Explanation
  - User-Friendly Interface
4. Exploring eBook Recommendations from Ap Chemistry Multiple Choice Explanation
  - Personalized Recommendations
  - Ap Chemistry Multiple Choice Explanation User Reviews and Ratings
  - Ap Chemistry Multiple Choice Explanation and Bestseller Lists

5. Accessing Ap Chemistry Multiple Choice Explanation Free and Paid eBooks
  - Ap Chemistry Multiple Choice Explanation Public Domain eBooks
  - Ap Chemistry Multiple Choice Explanation eBook Subscription Services
  - Ap Chemistry Multiple Choice Explanation Budget-Friendly Options
6. Navigating Ap Chemistry Multiple Choice Explanation eBook Formats
  - ePub, PDF, MOBI, and More
  - Ap Chemistry Multiple Choice Explanation Compatibility with Devices
  - Ap Chemistry Multiple Choice Explanation Enhanced eBook Features
7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Ap Chemistry Multiple Choice Explanation
  - Highlighting and Note-Taking Ap Chemistry Multiple Choice Explanation
  - Interactive Elements Ap Chemistry Multiple Choice Explanation
8. Staying Engaged with Ap Chemistry Multiple Choice Explanation
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Ap Chemistry Multiple Choice Explanation
9. Balancing eBooks and Physical Books Ap Chemistry Multiple Choice Explanation
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Ap Chemistry Multiple Choice Explanation
10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
11. Cultivating a Reading Routine Ap Chemistry Multiple Choice Explanation
  - Setting Reading Goals Ap Chemistry Multiple Choice Explanation
  - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Ap Chemistry Multiple Choice Explanation
  - Fact-Checking eBook Content of Ap Chemistry Multiple Choice Explanation
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

#### 14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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