

Questions no. 27 to 33 are short answer type questions.

**27.** Write one chemical equation each for the chemical reaction in which the following have taken place:

- (i) Change in colour
- (ii) Change in temperature
- (iii) Formation of precipitate

Mention colour change/temperature change (rise/fall)/compound precipitated along with the equation. [3]

**Solution:**

**(i) Change in colour:**

The reaction between lead nitrate solution and potassium iodide solution.



In this reaction colour changes from colourless to yellow due to formation of  $\text{PbI}_2$

**(ii) Change in temperature:**

The action of dilute sulphuric acid on calcium.



In this reaction, heat evolves.

**(iii) Formation of precipitate:**

The action of barium chloride on sodium sulphate.



Here,  $\text{BaSO}_4(\text{s})$  is a white precipitate.

**28.** (i) The pH of a sample of tomato juice is 4.6. How is this juice likely to be in taste? Give reason to justify your answer.

(ii) How do we differentiate between a strong acid and a weak base in terms of ion-formation in aqueous solutions?

(iii) The acid rain can make the survival of aquatic animals difficult. How? [3]

**Solution:**

(i) Tomato juice with a pH of 4.6 is considered slightly acidic. The taste of tomato juice at this pH level is likely to be tangy or slightly sour. This is because acidic substances tend to impart sourness or tanginess to foods and beverages.

(ii) Strong Acids : Ionize fully in water to give  $\text{H}^+$  ions in large numbers.

Weak Base : Ionize partially in water to give  $\text{OH}^-$  ions in small numbers.

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