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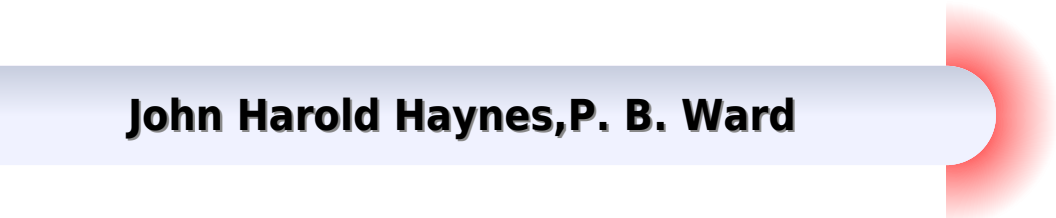
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John Harold Haynes,P. B. Ward



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 $2\text{H}_2\text{O} \rightarrow 2\text{H}_2 + \text{O}_2$
 $\text{Pb} + \text{OH}^- \rightarrow \text{PbCl}_2 + \text{H}_2\text{O}$
 $2\text{Al} + 3\text{K}_2\text{SO}_4 \rightarrow 6\text{KBr} + \text{Al}_2\text{SO}_4$
 $3\text{CH}_4 + 2\text{O}_2 \rightarrow \text{CO}_2 + 2\text{H}_2\text{O}$
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 $3\text{HBr} + \text{KHCO}_3 \rightarrow \text{H}_2\text{O} + \text{KBr} + \text{CO}_2$
 $\text{GaBr}_3 + \text{NaBr} \rightarrow \text{SNO}_4 + \text{NaSO}_3 + \text{Ga}_2\text{SO}_3 + 3\text{NF}_3 + \text{SnF}_2 + \text{N}_2\text{O}_3$ using the equation from problem 2 above answer the following

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 $3\text{HBr} + 1\text{KHCO}_3 \rightarrow 1\text{H}_2\text{O} + 1\text{KBr} + 1\text{CO}_2$
 $2\text{GaBr}_3 + 3\text{Na}_2\text{SO}_3 \rightarrow 1\text{Ga}_2\text{SO}_3 + 3\text{SO}_3 + 6\text{NaBr} + 3\text{SNO}_2 + 3\text{NF}_3 + 3\text{SnF}_2 + 1\text{N}_2\text{O}_3$ using the

following equation $2 \text{NaOH} + \text{H}_2\text{SO}_4 \rightarrow 2 \text{H}_2\text{O} + \text{Na}_2\text{SO}_4$

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 $\text{Br}_2 + \text{C}_6\text{H}_{10} \rightarrow 17 \text{O}_2 + 12 \text{CO}_2 + 10 \text{H}_2\text{O}$
 $\text{Br}_2 + \text{HBr} \rightarrow 1 \text{KHCO}_3 + 1 \text{H}_2\text{O} + 1 \text{KBr} + 1 \text{CO}_2$
 $\text{Br}_2 + 2 \text{GaBr}_3 \rightarrow 3 \text{Na}_2\text{SO}_3 + 1 \text{Ga}_2\text{SO}_3 + 3 \text{NaBr}$
 $\text{Br}_2 + 3 \text{SNO}_2 \rightarrow 2 \text{NF}_3 + 3 \text{SNF}_2 + 1 \text{N}_2\text{O}_3$
 using the following equation
 $\text{Br}_2 + 2 \text{NaOH} + \text{H}_2\text{SO}_4 \rightarrow 2 \text{H}_2\text{O} + \text{Na}_2\text{SO}_4$

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