

Lab 10: Calculations with a Chemical Structure

Objectives:

1. To derive an empirical formula from a chemical structure without performing stoichiometric calculations.
2. To calculate the number of moles of a solid from a weighing without performing stoichiometric calculations.
3. To determine the molar mass of a solid.
4. To determine the stoichiometric amount of a solid reagent that would be produced.
5. To compare the theoretical amount to the actual amount of a solid reagent and calculate the percent yield.

Reagents and Materials:

Reagent 1: Solid Chemical Reagent: Na_2CO_3 10.0 g

Reagent 2: Solid Chemical Reagent: NaCl 10.0 g

Reagent 3: Solid Reagent: Na_2CO_3 10.0 g

Reagent 4: Solid Reagent: NaCl 10.0 g

Reagent 5: Solid Reagent: Na_2CO_3 10.0 g

Reagent 6: Solid Reagent: NaCl 10.0 g

Reagent 7: Solid Reagent: Na_2CO_3 10.0 g

Procedure:

Reagents and Calculations:

1. Calculations:

$$\text{Moles of reagent 1} = \frac{\text{Mass of reagent 1}}{\text{Molar mass of reagent 1}} = \frac{10.0 \text{ g}}{105.99 \text{ g/mol}} = 0.0943 \text{ mol}$$

$$\text{Moles of reagent 2} = \frac{\text{Mass of reagent 2}}{\text{Molar mass of reagent 2}} = \frac{10.0 \text{ g}}{58.44 \text{ g/mol}} = 0.1711 \text{ mol}$$

$$\text{Moles of reagent 3} = \frac{\text{Mass of reagent 3}}{\text{Molar mass of reagent 3}} = \frac{10.0 \text{ g}}{105.99 \text{ g/mol}} = 0.0943 \text{ mol}$$

$$\text{Moles of reagent 4} = \frac{\text{Mass of reagent 4}}{\text{Molar mass of reagent 4}} = \frac{10.0 \text{ g}}{58.44 \text{ g/mol}} = 0.1711 \text{ mol}$$

2. Chemical Equations: $\text{Na}_2\text{CO}_3 + \text{NaCl} \rightarrow \text{Na}_2\text{CO}_3 + \text{NaCl}$

3. Stoichiometry:

$$\text{Moles of reagent 1} = \frac{\text{Mass of reagent 1}}{\text{Molar mass of reagent 1}} = \frac{10.0 \text{ g}}{105.99 \text{ g/mol}} = 0.0943 \text{ mol}$$

$$\text{Moles of reagent 2} = \frac{\text{Mass of reagent 2}}{\text{Molar mass of reagent 2}} = \frac{10.0 \text{ g}}{58.44 \text{ g/mol}} = 0.1711 \text{ mol}$$

$$\text{Moles of reagent 3} = \frac{\text{Mass of reagent 3}}{\text{Molar mass of reagent 3}} = \frac{10.0 \text{ g}}{105.99 \text{ g/mol}} = 0.0943 \text{ mol}$$

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