

SCH4U Unit 4 - Quiz – Equilibrium

Name: _____ Period: _____

Q1: (10)

Q2: (10)

Q3: (10)

Q4: (10)

Q5: (10)

Write the formulas for:

$$K_{eq} = \frac{[products]}{[reactants]} \quad ; \quad K_{eq} = \frac{P_{products}}{P_{reactants}}$$

$$K_{eq} = \frac{P_{CO_2}^2 P_{H_2O}}{P_{CH_4} P_{H_2O}^2} \quad ; \quad K_{eq} = \frac{P_{CO_2}^2 P_{H_2O}}{(P_{CH_4}) (P_{H_2O})^2}$$

gpt (aq) + h₂o (l) ⇌ gpt(s) + h₂g(g) ; gpt(s) ⇌ h₂g(g)

K_{eq} = K_{sp}/K_a

Q6: (10)

6. Hydrogen sulfide, H₂S, is a slightly poisonous gas. At 25°C, an equilibrium mixture contains 0.025 moles of hydrogen, H₂, 0.025 moles of hydrogen sulfide, with an equilibrium pressure of 0.400 atm. The reaction is:



If the volume of an 8.00 L container at 25°C where a concentration of H₂ is present at equilibrium:

- What is the concentration of H₂ at equilibrium?
 Answer: H₂ = 0.000625 moles/L
 Answer: H₂/H_{2S} = 0.000625 moles/L
 H₂ = 0.00125 mol
- Determine concentration of H_{2S} at equilibrium
- Answer:
 What other equilibria have associated formula constants at 25°C involving water?
- Calculate:

	K _{eq} = 1.0	K _{eq} = 1.0	K _{eq} = 1.0
Initial concentration (moles/L)			
Change in concentration (moles/L)			
Final concentration (moles/L)	0.00125	0.00125	0.00125

What is the final concentration?

Answers For Sch4u Unit 4

Tsfx



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