

## Atomic Structure

Atoms are made from protons, neutrons, and electrons. Protons and neutrons are known as nucleons.

Electrons have a negligible mass compared to protons and neutrons which both have a relative mass of 1. Protons and electrons have equal and opposite magnitudes of charge,  $1+$ ,  $1-$ .

**MASS NUMBER** – Total number of protons and neutrons in the nucleus

**ATOMIC NUMBER** – The number of protons in the nucleus of an atom.

Ions are formed when atoms lose or gain electrons.

**ISOTOPE** – An atom of the same element with a different number of neutrons.

Isotopes have the same chemical properties as they have the same electron configuration as one another. They have different physical properties to one another as this is due to the variance in mass.

The model of the atom has changed over time due to the increase in understanding:

- Dalton's model: Suggests that atoms are solid spheres and different spheres make up different elements.
- Thomson's Model: Atoms are a positive charge with negatively charged particles in them.
- Rutherford's Model: Small positive nucleus in the centre of the nucleus and electrons surrounding it, mostly empty space.
- Bohr's Model: Electrons exist in fixed orbitals and these have fixed energy.

**RELATIVE ATOMIC MASS** - Average mass of an atom of an element on a scale where  $1/12^{\text{th}}$  of  $\text{C}^{12}$  is exactly 1.

Steps of mass spectrometer:

- **Ionisation:**
  - o Electrospray  $\rightarrow$  Sample dissolved in a highly volatile solvent and injected through a hypodermic needle that is attached to the positive terminal of a high voltage power supply. The particles gain  $\text{H}^+$  and the solvent evaporates leaving gaseous positive ions.
  - o Electron impact  $\rightarrow$  High energy electrons are fired out of an electron gun which knocks out electrons from atoms.
- **Acceleration:** A negatively charged plate attracts electrons towards it and all ions are accelerated to the same kinetic energy.
- **Flight tube:** The ions drift through this region with the same speed as they left the electric field with.
- **Detection:** Ions hit a plate and gain an electron causing a current to flow, the size of this current can be measured to determine the abundance of each of the isotopes.

Mass spectrometry can be used to identify elements in a sample.

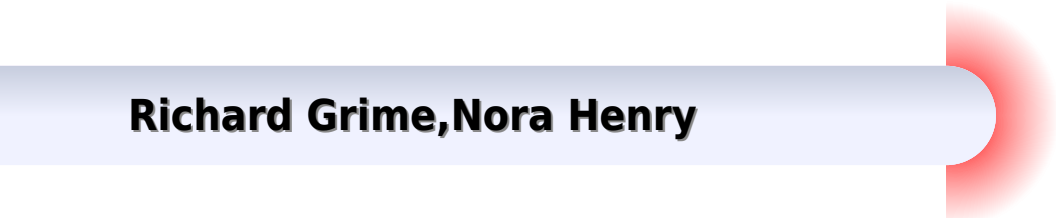


Figure 2: Diagram showing how a TOF mass spectrometer works.

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## **Table of Contents 20advanced Extension Aqa Chemistry Report**

1. Understanding the eBook 20advanced Extension Aqa Chemistry Report
  - The Rise of Digital Reading 20advanced Extension Aqa Chemistry Report
  - Advantages of eBooks Over Traditional Books
2. Identifying 20advanced Extension Aqa Chemistry Report
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an 20advanced Extension Aqa Chemistry Report
  - User-Friendly Interface
4. Exploring eBook Recommendations from 20advanced Extension Aqa Chemistry Report
  - Personalized Recommendations
  - 20advanced Extension Aqa Chemistry Report User Reviews and Ratings
  - 20advanced Extension Aqa Chemistry Report and Bestseller Lists

5. Accessing 20advanced Extension Aqa Chemistry Report Free and Paid eBooks
  - 20advanced Extension Aqa Chemistry Report Public Domain eBooks
  - 20advanced Extension Aqa Chemistry Report eBook Subscription Services
  - 20advanced Extension Aqa Chemistry Report Budget-Friendly Options
6. Navigating 20advanced Extension Aqa Chemistry Report eBook Formats
  - ePub, PDF, MOBI, and More
  - 20advanced Extension Aqa Chemistry Report Compatibility with Devices
  - 20advanced Extension Aqa Chemistry Report Enhanced eBook Features
7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of 20advanced Extension Aqa Chemistry Report
  - Highlighting and Note-Taking 20advanced Extension Aqa Chemistry Report
  - Interactive Elements 20advanced Extension Aqa Chemistry Report
8. Staying Engaged with 20advanced Extension Aqa Chemistry Report
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers 20advanced Extension Aqa Chemistry Report
9. Balancing eBooks and Physical Books 20advanced Extension Aqa Chemistry Report
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection 20advanced Extension Aqa Chemistry Report
10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
11. Cultivating a Reading Routine 20advanced Extension Aqa Chemistry Report
  - Setting Reading Goals 20advanced Extension Aqa Chemistry Report
  - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of 20advanced Extension Aqa Chemistry Report
  - Fact-Checking eBook Content of 20advanced Extension Aqa Chemistry Report
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

#### 14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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