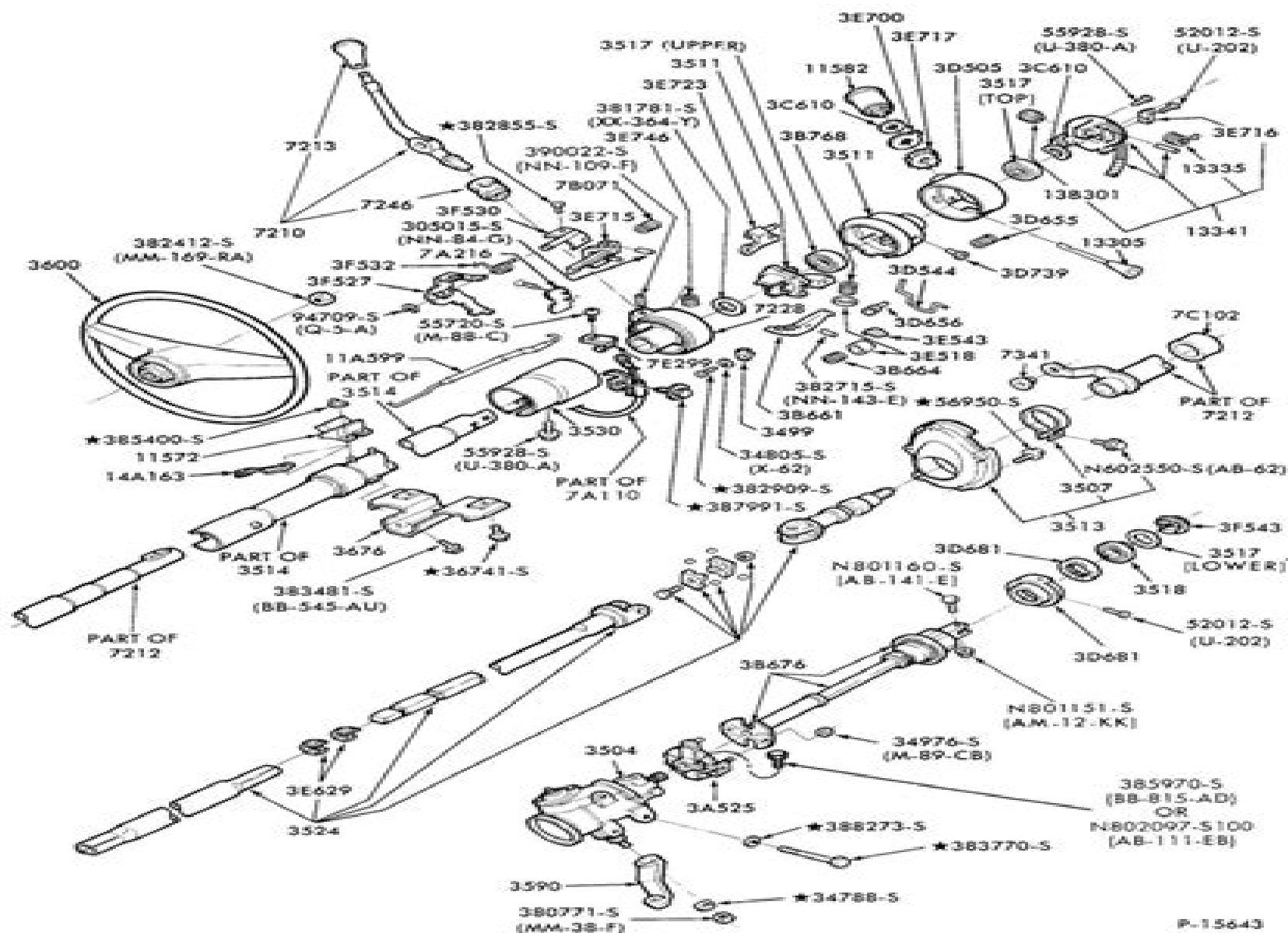


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web the  $\text{A}^-$  ion is a base because it is accepting a proton from the  $\text{H}_2\text{O}$  the reversible reaction contains two acids and two bases we think of them in pairs called conjugate pairs when the acid  $\text{HA}$  loses a proton it forms a base  $\text{A}^-$  when the base  $\text{A}^-$  accepts a proton back again it obviously reforms the acid  $\text{HA}$

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discovery by 1890 the swedish chemist svante arrhenius 1859 1927 was able to formulate the first useful theory of acids  
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electrons instead of proton transfer and specifically stated that an acid is a species that accepts

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principle is the acid base behavior of molecules in the general chemistry sequence the brønsted lowry theory is stressed  
because it lends itself well to studying equilibrium and kine