

1999 Advanced Placement Chemistry Exam

Section I: Multiple Choice

Time — 1 hour and 30 minutes

NO CALCULATORS MAY BE USED WITH SECTION I.

Note: For all questions, assume that the temperature is 298 K, the pressure is 1.00 atmosphere, and solutions are aqueous unless otherwise specified.

Throughout the test the following symbols have the definitions specified unless otherwise noted.

T = temperature
 P = pressure
 V = volume
 S = entropy
 H = enthalpy
 G = free energy
 R = molar gas constant
 n = number of moles
 M = molar

m = molal
 L, mL = liter(s), milliliter(s)
 g = gram(s)
 nm = nanometer(s)
 atm = atmosphere(s)
 J, kJ = joule(s), kilojoule(s)
 V = volt(s)
 mol = mole(s)

Part A

Directions: Each set of lettered choices below refers to the numbered statements immediately following it. Select the one lettered choice that best fits each statement and then fill in the corresponding oval on the answer sheet. A choice may be used once, more than once, or not at all in each set.

Questions 1–4 refer to the following types of energy.

- (A) Activation energy
- (B) Free energy
- (C) Ionization energy
- (D) Kinetic energy
- (E) Lattice energy

1. The energy required to convert a ground-state atom in the gas phase to a gaseous positive ion
2. The energy change that occurs in the conversion of an ionic solid to widely separated gaseous ions
3. The energy in a chemical or physical change that is available to do useful work
4. The energy required to form the transition state in a chemical reaction

Questions 5–8 refer to atoms for which the occupied atomic orbitals are shown below.

- (A) $1s \underline{\hspace{1cm}} 2s \underline{\hspace{1cm}}$
- (B) $1s \underline{\hspace{1cm}} 2s \underline{\hspace{1cm}}$
- (C) $1s \underline{\hspace{1cm}} 2s \underline{\hspace{1cm}} 2p \underline{\hspace{1cm}} \underline{\hspace{1cm}} \underline{\hspace{1cm}}$
- (D) $1s \underline{\hspace{1cm}} 2s \underline{\hspace{1cm}} 2p \underline{\hspace{1cm}} \underline{\hspace{1cm}} \underline{\hspace{1cm}}$
- (E) $[Ar] 4s \underline{\hspace{1cm}} 3d \underline{\hspace{1cm}} \underline{\hspace{1cm}} \underline{\hspace{1cm}} \underline{\hspace{1cm}}$

5. Represents an atom that is chemically unreactive
6. Represents an atom in an excited state
7. Represents an atom that has four valence electrons
8. Represents an atom of a transition metal

Questions 9–12 refer to aqueous solutions containing 1:1 mole ratios of the following pairs of substances. Assume all concentrations are 1 M .

Ap Chemistry Exam Multiple Choice Explanations

John T. Moore, Richard H. Langley



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Table of Contents Ap Chemistry Exam Multiple Choice Explanations

1. Understanding the eBook Ap Chemistry Exam Multiple Choice Explanations
 - The Rise of Digital Reading Ap Chemistry Exam Multiple Choice Explanations
 - Advantages of eBooks Over Traditional Books
2. Identifying Ap Chemistry Exam Multiple Choice Explanations
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Ap Chemistry Exam Multiple Choice Explanations
 - User-Friendly Interface
4. Exploring eBook Recommendations from Ap Chemistry Exam Multiple Choice Explanations
 - Personalized Recommendations
 - Ap Chemistry Exam Multiple Choice Explanations User Reviews and Ratings
 - Ap Chemistry Exam Multiple Choice Explanations and Bestseller Lists

5. Accessing Ap Chemistry Exam Multiple Choice Explanations Free and Paid eBooks
 - Ap Chemistry Exam Multiple Choice Explanations Public Domain eBooks
 - Ap Chemistry Exam Multiple Choice Explanations eBook Subscription Services
 - Ap Chemistry Exam Multiple Choice Explanations Budget-Friendly Options
6. Navigating Ap Chemistry Exam Multiple Choice Explanations eBook Formats
 - ePub, PDF, MOBI, and More
 - Ap Chemistry Exam Multiple Choice Explanations Compatibility with Devices
 - Ap Chemistry Exam Multiple Choice Explanations Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Ap Chemistry Exam Multiple Choice Explanations
 - Highlighting and Note-Taking Ap Chemistry Exam Multiple Choice Explanations
 - Interactive Elements Ap Chemistry Exam Multiple Choice Explanations
8. Staying Engaged with Ap Chemistry Exam Multiple Choice Explanations
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Ap Chemistry Exam Multiple Choice Explanations
9. Balancing eBooks and Physical Books Ap Chemistry Exam Multiple Choice Explanations
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Ap Chemistry Exam Multiple Choice Explanations
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Ap Chemistry Exam Multiple Choice Explanations
 - Setting Reading Goals Ap Chemistry Exam Multiple Choice Explanations
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Ap Chemistry Exam Multiple Choice Explanations
 - Fact-Checking eBook Content of Ap Chemistry Exam Multiple Choice Explanations
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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