

AP Chemistry Chapter 4 Practice Exam

D

1. The Claus reactions, shown below, are used to generate elemental sulfur from hydrogen sulfide.

How much sulfur (in grams) is produced from 21.5 grams of O_2 ?

A) 97.0 g

B) 14.4 g

C) 21.5 g

D) 43.1 g

E) none of these

$$\begin{array}{c} 21.5 \text{ g } \text{O}_2 \left| \frac{1 \text{ mol } \text{O}_2}{32.00 \text{ g } \text{O}_2} \right| \frac{2 \text{ SO}_2}{3 \text{ O}_2} \left| \frac{3 \text{ S}}{\text{SO}_2} \right| \frac{32.06 \text{ g S}}{1 \text{ mol S}} = 43.1 \text{ g S} \end{array}$$

B

2. Iron is produced from its ore by the reactions:

How many moles of $\text{O}_2(g)$ are needed to produce 4.6 moles of $\text{Fe}(s)$?A) 2.3 mole O_2 B) 3.5 mole O_2 C) 4.6 mole O_2 D) 6.9 mole O_2 E) 13.8 mole O_2

$$\begin{array}{c} 4.6 \text{ mol Fe}(s) \left| \frac{3 \text{ CO}}{2 \text{ Fe}} \right| \frac{1 \text{ O}_2}{2 \text{ CO}} = 3.45 \text{ mol O}_2 \end{array}$$

B

3. Suppose the reaction
- $\text{Ca}_3(\text{PO}_4)_2 + 3\text{H}_2\text{SO}_4 \rightarrow 3\text{CaSO}_4 + 2\text{H}_3\text{PO}_4$
- is carried out starting with 153 g of
- $\text{Ca}_3(\text{PO}_4)_2$

and 76.8 g of H_2SO_4 . How much phosphoric acid will be produced?

A) 76.7 g

B) 51.1 g

C) 229.8 g

D) 115.1 g

E) 96.7 g

$$\begin{array}{c} 153 \text{ g } \text{Ca}_3(\text{PO}_4)_2 \left| \frac{1 \text{ mol } \text{Ca}_3(\text{PO}_4)_2}{286.17 \text{ g}} \right| \frac{2 \text{ H}_3\text{PO}_4}{1 \text{ Ca}_3(\text{PO}_4)_2} = \\ 76.8 \text{ g } \text{H}_2\text{SO}_4 \left| \frac{1 \text{ mol } \text{H}_2\text{SO}_4}{98.08 \text{ g}} \right| \frac{2 \text{ H}_3\text{PO}_4}{3 \text{ H}_2\text{SO}_4} = \end{array}$$

B

- 4.
- SO_2
- reacts with
- H_2S
- as follows:

When 7.50 g of H_2S reacts with 12.75 g of SO_2 , which statement applies?

A) 6.38 g of sulfur are formed.

B) 10.6 g of sulfur are formed.C) 0.0216 moles of H_2S remain.D) 1.13 g of H_2S remain.E) SO_2 is the limiting reagent.

$$\begin{array}{c} 7.50 \text{ g } \text{H}_2\text{S} \left| \frac{1 \text{ mol } \text{H}_2\text{S}}{34.08 \text{ g}} \right| \frac{3 \text{ S}}{2 \text{ H}_2\text{S}} \left| \frac{32 \text{ g S}}{1 \text{ mol S}} \right| = 10.89 \end{array}$$

$$\begin{array}{c} 12.75 \text{ g } \text{SO}_2 \left| \frac{1 \text{ mol } \text{SO}_2}{64 \text{ g}} \right| \frac{3 \text{ S}}{1 \text{ SO}_2} \left| \frac{32 \text{ g S}}{1 \text{ mol S}} \right| = 19.13 \end{array}$$

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Neil D. Jespersen, Pamela Kerrigan



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