

Which of the following mixtures is a buffer with a pH of approximately 7?

- (A) A mixture of 100. mL of 0.1 M $\text{CH}_3\text{NH}_2\text{Cl}$ and 50. mL of 0.1 M NaOH
- (B) A mixture of 100. mL of 0.1 M $\text{HC}_2\text{H}_3\text{O}_2$ and 50. mL of 0.1 M NaOH
- (C) A mixture of 100. mL of 0.1 M $\text{NaC}_2\text{H}_3\text{O}_2$ and 100. mL of 0.1 M NaOH
- (D) A mixture of 100. mL of 0.1 M $\text{CH}_3\text{NH}_2\text{Cl}$ and 100. mL of 0.1 M CH_3NH_2



36. The Lewis electron-dot diagrams of the HClO_3 molecule and the HClO_4 molecule are shown above at the left and right, respectively. Which of the following statements identifies the stronger acid and correctly identifies a factor that contributes to its being the stronger acid?

- (A) $\text{HClO}_3(\text{aq})$ is the stronger acid because its molecules experience stronger London dispersion forces.
- (B) $\text{HClO}_3(\text{aq})$ is the stronger acid because the additional electronegative oxygen atom on the chlorine atom stabilizes the conjugate base.
- (C) $\text{HClO}_4(\text{aq})$ is the stronger acid because its



$$\Delta H^\circ = 178 \text{ kJ/mol}_{\text{rxn}}$$

40. The reaction system represented above is at equilibrium. Which of the following will decrease the amount of $\text{CaOH}(\text{s})$ in the system?

- (A) Increasing the volume of the reaction vessel at constant temperature
- (B) Lowering the temperature of the system
- (C) Removing some $\text{CO}_2(\text{g})$ at constant temperature
- (D) Removing some $\text{CaCO}_3(\text{s})$ at constant temperature

Acid Solution	Volume of NaOH Added (mL)
A	40
B	75
C	115
D	200

$$M_A V_A = M_B V_B$$

41. To maximize the yield in a certain manufacturing process, a solution of a weak monoprotic acid that has a concentration between 0.20 M and 0.30 M is required. Four 100. mL samples of the acid at different concentrations are each titrated with a 0.20 M NaOH solution. The volume of NaOH needed to reach the end point for each sample is shown in the table above. Which solution is the

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M Carnoy



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