

Balancing redox equations

Example

Balance the equation showing the oxidation of Fe^{2+} ions to Fe^{3+} ions by dichromate ions ($\text{Cr}_2\text{O}_7^{2-}$) in an acidic medium



Step 1: Identify oxidising and reducing agents and write half reactions



$$2\text{Cr} = 12$$

$$\text{Cr} = +6$$



Iron loses one electron, i.e. it is *oxidised*. It donates one electron to dichromate, i.e. it is the *reducing agent*.

Half equation: Oxidation reaction: $\text{Fe}^{2+} \rightarrow \text{Fe}^{3+} + \text{e}^-$



Chromium gains three electrons, i.e. it is *reduced*. It gains its electrons from iron, i.e. it is the *oxidising agent*.

Half equation: Reduction reaction: $\text{Cr}_2^{6+} + 3\text{e}^- \rightarrow \text{Cr}^{3+}$

Analyzing Oxidation Reduction Reactions Answers

Chunlong Zhang



Analyzing Oxidation Reduction Reactions Answers:

Enjoying the Song of Term: An Psychological Symphony within **Analyzing Oxidation Reduction Reactions Answers**

In a global eaten by displays and the ceaseless chatter of immediate connection, the melodic beauty and mental symphony created by the written term often diminish in to the back ground, eclipsed by the relentless noise and distractions that permeate our lives. Nevertheless, nestled within the pages of **Analyzing Oxidation Reduction Reactions Answers** a marvelous literary prize filled with fresh feelings, lies an immersive symphony waiting to be embraced. Crafted by a wonderful composer of language, this charming masterpiece conducts visitors on a mental trip, well unraveling the hidden tunes and profound affect resonating within each carefully crafted phrase. Within the depths of this touching assessment, we can examine the book is central harmonies, analyze its enthralling writing model, and submit ourselves to the profound resonance that echoes in the depths of readers souls.

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