

# Gravimetric Analysis of Chloride in Solution

Experiment 10  
Lab 10  
Date: \_\_\_\_\_

## Introduction:

This experiment involves the gravimetric analysis of a sample solution containing chloride ions. The sample solution is a mixture of sodium chloride and other soluble salts. The sample is weighed and then dissolved in water. The solution is then treated with silver nitrate, which reacts with the chloride ions to form a precipitate of silver chloride. The precipitate is then filtered, washed, and dried. The mass of the precipitate is then measured, and the mass of chloride in the original sample is calculated. The experiment is performed in a fume hood to avoid exposure to silver nitrate, which is a strong oxidizing agent and can cause skin irritation. The experiment is also performed in a clean, dry environment to avoid contamination of the sample.

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## Procedure:

The procedure for this experiment is as follows: 1. Weigh a sample of the solution (approximately 0.5 g) into a clean, dry beaker. 2. Add approximately 10 mL of distilled water to the beaker. 3. Stir the solution with a glass rod. 4. Add approximately 1 mL of 10% silver nitrate solution to the beaker. 5. Stir the solution with a glass rod. 6. Filter the solution through a filter paper into a clean, dry beaker. 7. Wash the filter paper with distilled water. 8. Dry the filter paper in a oven at 100°C for 1 hour. 9. Weigh the dried filter paper and the precipitate. 10. Calculate the mass of chloride in the original sample.

## Data Analysis:

|                            | Sample 1 | Sample 2 |
|----------------------------|----------|----------|
| Mass of sample             | 0.5 g    | 0.5 g    |
| Mass of precipitate        | 0.2 g    | 0.2 g    |
| Mass of chloride in sample | 0.1 g    | 0.1 g    |

The mass of chloride in the sample is calculated as follows:  $\text{Mass of chloride} = \text{Mass of precipitate} \times \frac{\text{Molar mass of chloride}}{\text{Molar mass of silver chloride}}$

For Sample 1:  $\text{Mass of chloride} = 0.2 \text{ g} \times \frac{35.5 \text{ g/mol}}{143.5 \text{ g/mol}} = 0.1 \text{ g}$

For Sample 2:  $\text{Mass of chloride} = 0.2 \text{ g} \times \frac{35.5 \text{ g/mol}}{143.5 \text{ g/mol}} = 0.1 \text{ g}$

# Analysis Of Chloride Experiment 7 Lab Report

**Judd E. Hollander**



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