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Chemistry Paper 3A WASSCE (SC), 2022

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Question 1A

A is an aqueous solution of iodine. **B** is 0.100 mol dm⁻³ sodium trioxothiosulphate (II).

- a. Put **B** into the burette. Pipette 20.0 cm³ or 25.0 cm³ of **A** into a conical flask. Add **E** from the burette until the reddish-brown colour fades to pale yellow, then add a few drops of starch indicator to obtain a dark blue solution. Continue adding **B** slowly from the burette until one drop of **B** causes the blue colour to disappear, leaving a colourless solution.

Repeat the titration to obtain concordant titre values.

Tabulate your results and calculate the average volume of **B** used.

The equation for the reaction is:



- b. From your results and the information provided, calculate the:

- i. Concentration in mol dm⁻³ of iodine in **A**;
- ii. mass in grammes of iodine in 1.00 dm³ of **A**.

[I=127.0]

[17 marks]

Observation

Majority of the candidates responded to this question and their performance was above average.

In part (a), majority of the candidates carried out the titration experiment correctly;

In part (b), majority of the candidates calculated the concentration in mol dm⁻³ of iodine in **A** and mass of iodine in 1.00 dm³ of **A**.

The expected answers include:

2014 Waec Chemistry Paper 3 Answer

Misra, Sanjay, Adewumi, Adewole



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