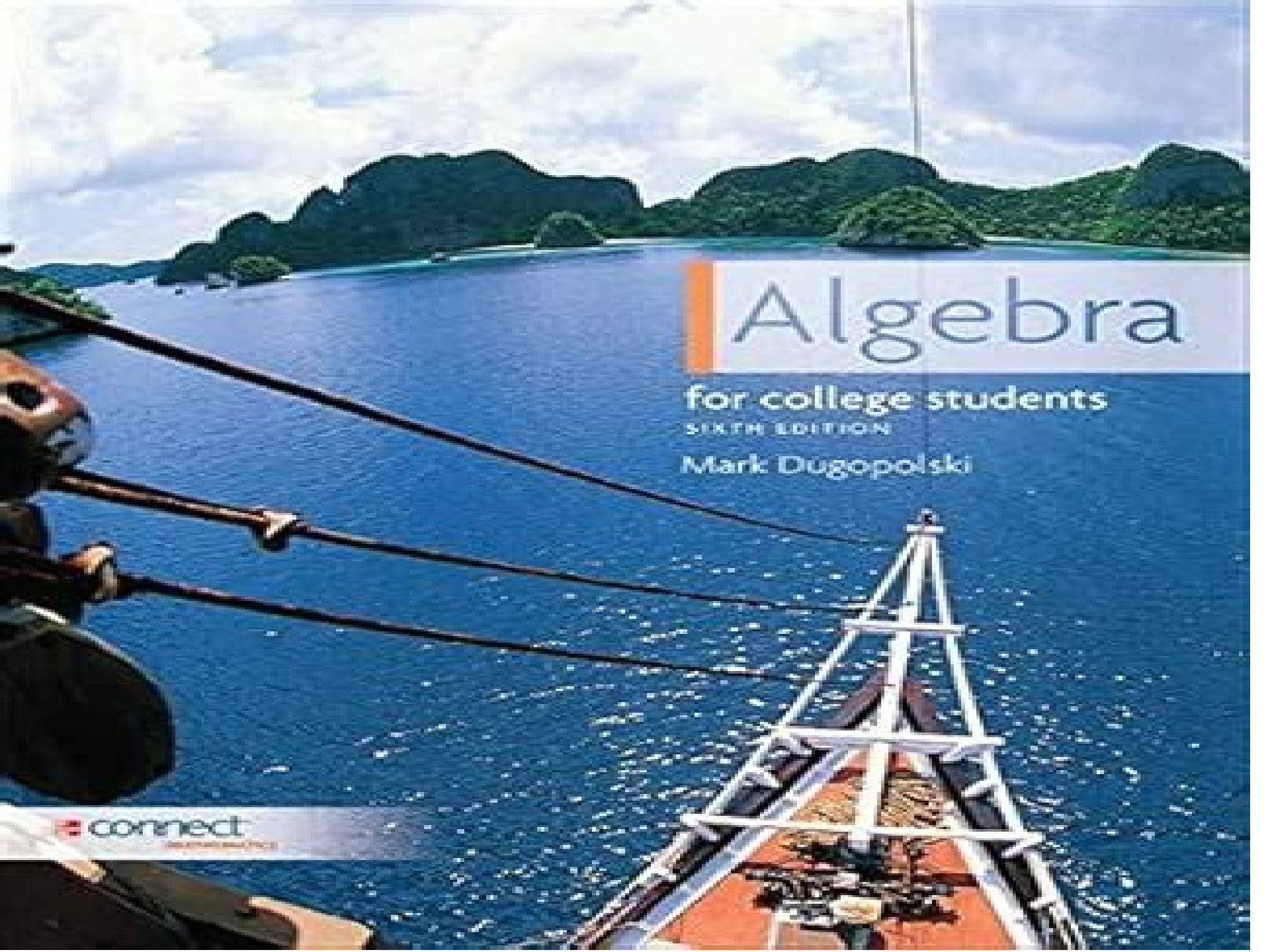




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web identify the changes in oxidation states and write the oxidation half reaction and the reduction half reaction and for each

half reaction balance the atom undergoing redox changes if necessary add the number of electrons that correspond to the

change in oxidation state

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reaction a redox reaction explain your answer  $2\text{NaCl} + \text{Pb} \rightarrow \text{PbCl}_2 + 2\text{Na}$

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web 1 define each remember oil rig oxidation is loss of e reduction is gain of e a oxidation b reduction c oxidizing agent d

reducing agent 2 write half reactions for each of the following atoms or ions label each as oxidation or reduction use the

activity series chart a Al b Ba<sup>2+</sup> c Br<sub>2</sub> d Ca e Ga<sup>3+</sup> f H<sub>2</sub> g H

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web during oxidation and reduction the oxidation numbers of atoms change if an atom is oxidized its oxidation number

increases ie it becomes more ve or less ve if an atom is reduced its oxidation number decreases ie it becomes less ve or more

ve

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web you should try to answer the questions without referring to your textbook if you get stuck try asking another group for

help 1 identify the oxidation state of each element in the following a Na<sub>2</sub>SO<sub>4</sub> b MnO<sub>4</sub><sup>-</sup> c KClO<sub>3</sub> d CH<sub>3</sub>OH 2 identify the

oxidation state of nitrogen in the following a NO<sub>3</sub><sup>-</sup> b NO<sub>2</sub> c NO<sub>2</sub>

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web an oxidation reduction redox reaction is one in which one or more atoms change oxidation numbers oxidation occurs

when an atom s oxidation state becomes more positive indicating that electrons have been lost reduction occurs when an

atom s oxidation state becomes less positive indicating that electrons have been gained

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and loss in oxygen reduction and oxidation by gain and loss of electrons reduction and oxidation by gain and loss of electron numbers redox of complex ions oxidising and reducing agents suited for student in y10 and y11

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web oxidation reduction reactions commonly known as redox reactions are reactions that involve the transfer of electrons from one species to another the species that loses electrons is said to be oxidized while the species that gains electrons is said to be reduced

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web describe redox reactions in terms of electron transfer recall that oxidation involves loss of electrons and reduction involves gain of electrons represent simple redox reactions using half equations identify oxidising and reducing agents for simple redox reactions understand the reactivity of metals and metal displacement reaction

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web what is an oxidation reduction or redox reaction this is a reaction in which electrons are transferred between reactants what do the following terms mean oxidation oxidation is a loss of electrons reduction reduction is a gain of electrons what is a helpful way to remember this

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web write the oxidation and reduction reactions for each redox reaction the first one is done for you 23  $\text{Fe} + 2\text{Co} \rightarrow \text{Fe}^{2+} + 2\text{Co}^{2+}$  oxidation  $\text{Co} \rightarrow \text{Co}^{2+} + 2\text{e}^{-}$  reduction  $\text{Fe}^{2+} + 2\text{e}^{-} \rightarrow \text{Fe}$  24 3  $\text{Ag} + \text{Ni} \rightarrow 3\text{Ag} + \text{Ni}^{2+}$  oxidation reduction 25  $\text{Cu} + \text{Pb} \rightarrow \text{Cu}^{2+} + \text{Pb}$  oxidation reduction 26  $\text{O}_2 \rightarrow 2\text{O}^{2-}$

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web in each of the following redox reactions indicate the elements that undergo changes in oxidation number and the magnitude of change then identify the element oxidized and reduced as well as the oxidizing agent and reducing agent a  $\text{I}_2 + 5\text{CO} \rightarrow 2\text{IO}_5 + 5\text{CO}_2$  b  $\text{I}_2 + 5\text{CO} \rightarrow 2\text{IO}_5 + 5\text{CO}_2$

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web chemical reactions in which there is a transfer of electrons from one substance to another are known as oxidation reduction reactions or redox reactions in this unit you will examine the oxidation reduction process and use the oxidation state and oxidation number concepts to identify redox reactions and to keep track of electrons

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web jun 28 2010 answers i determine what is oxidized and what in each reaction identify the oxidizing agent and the reducing agent also  
 $2\text{Sr} + \text{O}_2 \rightarrow 2\text{SrO}$   $\text{Sr}^0$  to  $\text{Sr}^2$  oxidized reducing agent  $\text{O}^0$  to  $\text{O}^2$  reduced ox ag  
 $2\text{Li} + \text{S} \rightarrow \text{Li}_2\text{S}$   $\text{Li}^0$  to  $\text{Li}^1$  oxidized red ag  
 $\text{S}^0$  to  $\text{S}^2$  reduced ox ag  
 $3\text{C} + 2\text{Br}_2 \rightarrow \text{C}_2\text{Br}_4 + \text{C}$