

Gravimetric Analysis of Chloride in Solution

Experiment 10
Lab 10
Date: _____

Introduction

This experiment involves the gravimetric analysis of a sample solution containing chloride ions. The sample solution is a mixture of sodium chloride and sodium nitrate. The sample is weighed and then dissolved in water. The solution is then treated with silver nitrate, which reacts with the chloride ions to form a precipitate of silver chloride. The precipitate is then filtered, dried, and weighed. The mass of the precipitate is used to calculate the mass of chloride in the original sample. The experiment is performed in a fume hood to avoid exposure to silver nitrate, which is a strong oxidizing agent and can cause skin irritation. The experiment is also performed in a clean, dry environment to avoid contamination of the sample.

The sample is weighed and then dissolved in water. The solution is then treated with silver nitrate, which reacts with the chloride ions to form a precipitate of silver chloride. The precipitate is then filtered, dried, and weighed. The mass of the precipitate is used to calculate the mass of chloride in the original sample. The experiment is performed in a fume hood to avoid exposure to silver nitrate, which is a strong oxidizing agent and can cause skin irritation. The experiment is also performed in a clean, dry environment to avoid contamination of the sample.

Procedure

The procedure involves the following steps: 1. Weighing the sample. 2. Dissolving the sample in water. 3. Adding silver nitrate to the solution. 4. Filtering the precipitate. 5. Drying the precipitate. 6. Weighing the dried precipitate. 7. Calculating the mass of chloride in the sample.

Data Analysis

	Sample 1	Sample 2
Mass of sample	0.500 g	0.500 g
Mass of precipitate	0.850 g	0.850 g
Mass of chloride in sample	0.350 g	0.350 g

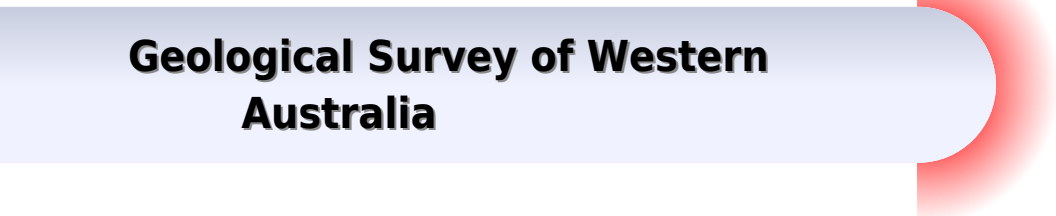
The mass of chloride in the sample is calculated using the following equation:

$$\text{Mass of chloride} = \text{Mass of precipitate} \times \frac{\text{Molar mass of chloride}}{\text{Molar mass of silver chloride}}$$

The mass of chloride in the sample is calculated using the following equation:

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