

Electrostatic Potential Lab

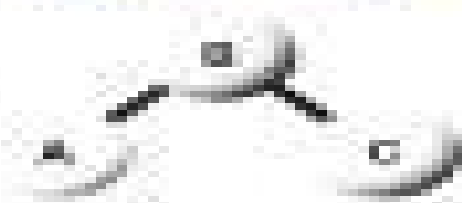
Name _____ Per _____

Circle the molecule below: Circle the ring carbon in the structure.

Part A: Three atom investigation

1. Click the "Keep the box" button adjacent to the three atoms in the electrostatic potential map and the bond charges plot.
2. Change the electronegativity of **A** and **B** to **1.0** and **2.0** respectively. The electronegativity of atom **C** is 2.5. Observe the color, partial charges and bond structure. Write the following information: the electronegativity of atom **B** is _____, the partial charges are _____, and the bond structure is _____.
3. Now put the electronegativity of atom **B** to the same electronegativity as the electronegativity of **A** and observe the results. Fill in the following information:
 - a. The polarity arrow points to the _____ electronegative atom.
 - b. The partial positive charge is always on the _____ electronegative atom.
 - c. The larger the electronegativity difference is, the _____ the bond structure.

Part B: Three atom investigation

1. Change the electronegativity of atom **A** to 2.5 and the electronegativity of atom **B** to 1.0. Circle the ring carbon in the electrostatic potential map and the bond charges plot.
 2. Change the electronegativity of the three atoms in the diagram.
 3. In order to make the electronegativity difference the same in the diagram.
 4. Change the partial charge distribution position in the diagram.
 5. Circle the electronegativity. Write the other electronegativity the molecule is represented.
- 
1. Click the "Keep the box" button adjacent to the three atoms in the electrostatic potential map and the bond charges plot.
 2. Change the electronegativity of atom **A** to 2.5 and the electronegativity of atom **B** to 1.0. Circle the ring carbon in the electrostatic potential map and the bond charges plot.
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 6. Circle the electronegativity. Write the other electronegativity the molecule is represented.
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Answer Of Molecular Polarity Phet Lab

Matthew Schnurr, Anna MacLeod



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