

Name: Samantha Norris Date: april 31 2020

Student Exploration: Calorimetry Lab

Vocabulary: calorie, calorimeter, joule, specific heat capacity

Gizmo Warm-up

A calorimeter is an insulated container filled with a liquid, usually water. When a hot object is placed in the calorimeter, heat energy is transferred from the object to the water and the water heats up.

Calorimeters can be used to find a substance's **specific heat capacity**. You will use the *Calorimetry Lab* Gizmo to determine the specific heat capacities of various substances.



- On the SIMULATION pane, select **Copper**. Use the slider to set its **Mass** to 200 g. Set the **Water mass** to 200 g. Check that the **Water temp** is set to 30.0 °C and the copper's **Temp** is 90 °C. Select the GRAPH tab, and click **Play** (▶).
 - What was the **Final temperature** of the copper and the water?
34.96
 - How much did the temperature of the copper change?
55.04
 - How much did the temperature of the water change?
4.96
- Specific heat capacity can be described as a substance's resistance to temperature changes. Which substance has a greater specific heat capacity, copper or water? Explain.

Water has a greater specific heat because water resists temperature better than copper.

Answer Sheet For Calorimetry Lab

James R. Sills

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the name bar.

Answer Sheet For Calorimetry Lab:

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content importance selection and classification of subject matter with reference to age sex and differently abled pupils
Integrated programme for boys and girls Teaching aids Time table Concepts credit system for various subject courses theory
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other means and Immunization Psychosomatic disorders sedentary life style diseases causes symptoms and prevention
Obesity related health problems Body weight control and its significance on health Role of exercise dieting and combination
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