

Name: _____ Course & Year: _____ Score: _____
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Activity No. 4
 WORKSHEET

CALORIMETRY AND SPECIFIC HEAT

Data, results, and calculations: Complete the table below, assume that the metal used is iron

(a) Weight of the calorimeter (inner vessel)	9.15 g
(b) Weight of Calorimeter + Water	36.70 g
(c) Weight of water, (b) - (a)	
(d) Initial temperature of water, (t) _w	28 °C
(e) Initial temperature of aluminum calorimeter, (t) _{cal}	
(f) Initial temperature of metal (t) _{metal}	98.0 °C
(g) Weight of metal sample, m _{metal}	7.5 g
(h) Final temperature of water, (t) _w	32.5 °C
(i) Final temperature of calorimeter, (t) _{cal}	
(j) Final temperature of metal	
(k) Heat absorbed by the calorimeter, q _{cal}	
(l) Heat absorbed by the water, q _w	
(m) Experimental specific heat of metal sample	
(n) Theoretical specific heat of metal sample	0.460 cal/g°C
(o) Percent Error	%

GUIDE QUESTIONS

1	What is the purpose of stirring the contents of the calorimeter before the final temperature is taken?	
2	Why is it required to tap the hot metal with a tissue paper before placing the hot metal in the calorimeter?	
3	Why should the transferring of the hot metal from the beaker to the calorimeter be done quickly?	

Answer Sheet For Calorimetry Specific Heat

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