

SCH4U Unit 4 - Quiz – Equilibrium

Player: _____

Period: _____

120 130 140 150

160 170

180

190 200

210

Student Number(s): _____

$$K_{eq} = \frac{[products]}{[reactants]} \quad ; \quad K_{eq} = \frac{[C]^c [D]^d}{[A]^a [B]^b}$$

$$K_{eq} = \frac{P_{CO_2}^2 P_{H_2O}}{P_{CH_4}^2 P_{O_2}} \quad ; \quad K_{eq} = \frac{P_{CO_2}^2 P_{H_2O}}{(P_{CH_4})^2 (P_{O_2})^2}$$

gpt (g/mol) x mol / gpt (g) = mol (gpt) / mol (gpt) = mol (gpt)

$1.2 \times 10^{-2} \times 16.04$

Question 1:

1. Hydrogen sulfide, $H_2S(g)$, is a slightly poisonous gas. At 1000°C, an equilibrium mixture contains 0.020 moles of hydrogen, $H_2(g)$, 0.010 moles of hydrogen sulfide, with an equilibrium pressure of 0.600 atm. The reaction is:



If the volume of an 8.00 L container at 1000°C where a concentration of $H_2(g)$ is present at equilibrium:

- What is the concentration of $H_2(g)$ at equilibrium?
 Answer: $[H_2(g)] = 0.020 \text{ mol/L}$
 Answer: $[H_2(g)] = 0.010 \text{ mol/L}$
 Answer: $[H_2(g)] = 0.005 \text{ mol/L}$
- Determine concentration of $H_2S(g)$ at equilibrium.
- Answer:
 What other species have chemical formula consisting of 2 hydrogen atoms alone?
- Calculate:

	$2H_2(g)$	$S_2(g)$	$2H_2S(g)$
Initial (moles present)			
Change (moles present)			
Final (moles present)	0.020	0.010	0.010

What is the final concentration?

Answers For Sch4u Unit 4

PT Brinkman

A red circular graphic with a gradient, appearing as a stylized arrow or a partial circle, located to the right of the PT Brinkman text.

Answers For Sch4u Unit 4:

4th Grade Science 2nd Edition Unit 4 BrightThinker,2026-08
Maths - Question Bank Tsfx,2020-06 *Open court grade 4 unit 4 ,*

TSFX - Exam Essentials - Unit 4 Specialist

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