

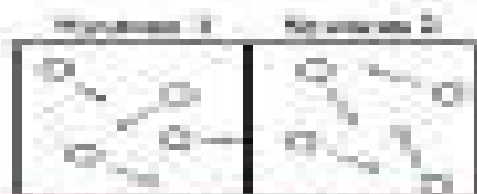
Unit 6: Thermochemistry**6.1: Heat Transfer and Thermal Equilibrium****6.2: Heat Capacity and Calorimetry****6.3: Entropy and Spontaneous Changes**

2. Use the Boltzmann-probability distribution curves below to predict temperatures of the mixtures of particles.



Temperature: is a measure of the average kinetic energy of the particles in a system. The kinetic energy for particles moving with v is $\frac{1}{2}mv^2$, so the more of particles that move fast enough, the more kinetic energy we will see. Temperature is a scalar. The Boltzmann-probability curves show that the average velocity we will see for the distribution is increasing from $T1$ to $T3$, so the average temperature is increasing.

3. Sketch what you expect to represent the relative velocities of the particles in the following two gases, samples of the same gas.



- a. In System 1, I see a higher or lower temperature than System 2? Justify your answer.
 System 1 is at a higher temperature than System 2. Temperature is a measure of the average KE of the particles in the system. In System 1, the particles are more spread out, so the average KE of the particles is higher. In System 2, the particles are more crowded, so the average KE of the particles is lower. Therefore, the temperature of System 1 is higher than the temperature of System 2.
- b. In which direction will thermal energy be transferred? Justify your answer in terms of molecular collisions.
 Thermal energy will be transferred from System 1 to System 2. The particles in System 1 are moving faster than the particles in System 2. When the particles from the two systems collide, the particles from System 1 will transfer some of their kinetic energy to the particles from System 2. This will continue until the temperatures of the two systems are equal.

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