

Stoichiometry

Solve the following problems:

1. Four moles of iron (II) sulfide added to seven moles of oxygen, produced two moles of iron (III) oxide and four moles of sulfur dioxide. (20pts)

a) Translate the problem to chemical equation.



Reactant side

4 Fe atoms

4 S atoms

7(2) = 14 O atoms

Product side

4 Fe atoms

4 S atoms

2(3) + 4(2) = 14 O atoms

- b) How many grams of oxygen will react with 2.0 grams of iron (II) sulfide?

Given:

m FeS = 2.0 g

Req'd. = m O₂

Solution:

molar mass of FeS = 88 g/mol

mole ratio = 4 mol FeS: 7 mol O₂

molar mass of O₂ = 32 g/mol

$$m_{O_2} = 2.0 \cancel{g_{FeS}} \left(\frac{1 \cancel{mol_{FeS}}}{88 \cancel{g_{FeS}}} \right) \left(\frac{7 \cancel{mol_{O_2}}}{4 \cancel{mol_{FeS}}} \right) \left(\frac{32 \cancel{g_{O_2}}}{1 \cancel{mol_{O_2}}} \right) = 1.3 g_{O_2}$$

2. Rusting involves the reaction of iron, Fe with oxygen, O₂ to form iron (III) oxide, Fe₂O₃. (20pts)

a) Write the balanced chemical equation.



Reactant side

4 Fe atoms

3(2) = 6 O atoms

Product side

4 Fe atoms

2(3) = 6 O atoms

- b) What are the mole ratios of the product to each reactant and of the reactants to each other?

Reactant to each other:

4 mol Fe: 3 mol O₂

Product to each reactant:

2 mol Fe₂O₃: 4 mol Fe

Answer Key To Stoichiometry Homework Problems

H. Scott Fogler



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web dərslərlər c programlaşdırma dili operatorlar c programlaşdırma dili operatorlar dərslər haqqında addımlar müəllif
operatorlar Əvvəlki paragrafda biz mənimsətmə operatoru ilə tanış olduq bu paragrafda isə şərt if switch və dövr for while do
while operatorları ilə tanış olacağıq Şərt operatorları if switch