

Chemistry Unit 8 Worksheet 4

Samples of Every Kind of Problem

On a separate sheet of paper, write a complete solution to each of the problems below. Follow the procedure outlined in class. Be sure to circle your final answer.

1. Calculate the number of moles of potassium chlorate, KClO_3 (s), that must decompose to produce potassium chloride, KCl (s), and 1.8 moles of oxygen gas.

Equation:	$2 \text{KClO}_3 \rightarrow 2 \text{KCl} + 3 \text{O}_2$			1.2 moles KClO_3 needed
Before:	1.2	0	0	
Change:	-1.2	+1.2	+1.8	
After:	0	1.2	1.8	

2. In a single displacement reaction, magnesium metal reacts with hydrochloric acid to produce magnesium chloride and hydrogen gas. How many moles of hydrochloric acid are needed to completely react with 2.43 g of magnesium?

Equation:	$\text{Mg} + 2 \text{HCl} \rightarrow \text{MgCl}_2 + \text{H}_2$				0.200 moles HCl needed
Before:	0.100	?	0	0	
Change:	-0.100	-0.200 +	0.100 +	0.100	
After:	0	?	0.100	0.100	

3. Ethane, C_2H_6 reacts with oxygen gas to produce carbon dioxide gas and water vapor. What mass of oxygen gas is required to react with 2.20 moles of ethane?

Equation:	$2 \text{C}_2\text{H}_6 + 7 \text{O}_2 \rightarrow 4 \text{CO}_2 + 6 \text{H}_2\text{O}$			
Before:	2.20	xs	0	0
Change:	-2.20	-7.70	+4.40	+6.60
After:	0	xs	4.40	6.60

$$\text{O}_2 \text{ needed: } 7.70 \text{ mole} \times \frac{32.0 \text{ g}}{1 \text{ mole}} = 246 \text{ g O}_2 \text{ required}$$

4. Determine the mass of sodium nitrate produced when 0.73 g of nickel (II) nitrate reacts with sodium hydroxide according to the following equation:



$$0.73 \text{ g} \times \frac{1 \text{ mole}}{183 \text{ g}} = .0040 \text{ mole}$$

Equation:	$\text{Ni}(\text{NO}_3)_2 + 2 \text{NaOH} \rightarrow \text{Ni}(\text{OH})_2 + 2 \text{NaNO}_3$			
Before:	0.0040	XS	0	0
Change:	-0.0040	-0.0080	+0.0040	+0.0080
After:	0	XS	0.0040	0.0080

$$\text{NaNO}_3 \text{ produced: } 0.0080 \text{ mole} \times \frac{85.0 \text{ g}}{1 \text{ mole}} = 0.68 \text{ g}$$

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Table of Contents Answer Sheet Chemistry Unit 8 Worksheet 4

1. Understanding the eBook Answer Sheet Chemistry Unit 8 Worksheet 4
 - The Rise of Digital Reading Answer Sheet Chemistry Unit 8 Worksheet 4
 - Advantages of eBooks Over Traditional Books
2. Identifying Answer Sheet Chemistry Unit 8 Worksheet 4
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Answer Sheet Chemistry Unit 8 Worksheet 4
 - User-Friendly Interface
4. Exploring eBook Recommendations from Answer Sheet Chemistry Unit 8 Worksheet 4
 - Personalized Recommendations
 - Answer Sheet Chemistry Unit 8 Worksheet 4 User Reviews and Ratings
 - Answer Sheet Chemistry Unit 8 Worksheet 4 and Bestseller Lists
5. Accessing Answer Sheet Chemistry Unit 8 Worksheet 4 Free and Paid eBooks
 - Answer Sheet Chemistry Unit 8 Worksheet 4 Public Domain eBooks
 - Answer Sheet Chemistry Unit 8 Worksheet 4 eBook Subscription Services
 - Answer Sheet Chemistry Unit 8 Worksheet 4 Budget-Friendly Options
6. Navigating Answer Sheet Chemistry Unit 8 Worksheet 4 eBook Formats
 - ePub, PDF, MOBI, and More
 - Answer Sheet Chemistry Unit 8 Worksheet 4 Compatibility with Devices
 - Answer Sheet Chemistry Unit 8 Worksheet 4 Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Answer Sheet Chemistry Unit 8 Worksheet 4
 - Highlighting and Note-Taking Answer Sheet Chemistry Unit 8 Worksheet 4
 - Interactive Elements Answer Sheet Chemistry Unit 8 Worksheet 4
8. Staying Engaged with Answer Sheet Chemistry Unit 8 Worksheet 4

- Joining Online Reading Communities
- Participating in Virtual Book Clubs
- Following Authors and Publishers Answer Sheet Chemistry Unit 8 Worksheet 4
- 9. Balancing eBooks and Physical Books Answer Sheet Chemistry Unit 8 Worksheet 4
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Answer Sheet Chemistry Unit 8 Worksheet 4
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Answer Sheet Chemistry Unit 8 Worksheet 4
 - Setting Reading Goals Answer Sheet Chemistry Unit 8 Worksheet 4
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Answer Sheet Chemistry Unit 8 Worksheet 4
 - Fact-Checking eBook Content of Answer Sheet Chemistry Unit 8 Worksheet 4
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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