

Lab 8B: Calculations with a Chemical Structure

Objectives:

1. To derive an empirical formula from a chemical structure without performing stoichiometric calculations.
2. To calculate the number of moles of a solid from a weighing measurement in the lab.
3. To determine the molar mass of a solid.
4. To determine the stoichiometric amount of a solid reagent that would be produced.
5. To compare the theoretical amount to the actual amount of substance formed and calculate the percent yield.

Reagents and Materials:

Reagent 1: Sodium Chloride Solution: 0.5 M and

Reagent 2: Calcium Chloride Solution: 0.5 M and

Reagent 3: Silver Nitrate Solution: 0.5 M and

Reagent 4: Silver Nitrate Solution: 0.5 M and

Reagent 5: Silver Nitrate Solution: 0.5 M and

Reagent 6: Silver Nitrate Solution: 0.5 M and

Reagent 7: Silver Nitrate Solution: 0.5 M and

Procedure:

Reagents and Materials:

1. Reagent 1:

2. Reagent 2:

3. Reagent 3:

4. Reagent 4:

5. Reagent 5:

6. Reagent 6:

7. Reagent 7:

8. Reagent 8:

9. Reagent 9:

10. Reagent 10:

Answers For Lab 6b Heath Chemistry

Arthur James Wells



Answers For Lab 6b Heath Chemistry:

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