

CHAPTER 10: Reactions

Reaction Energy

SECTION 1

Write a balanced chemical equation for the following reactions in the space provided.

1. For elements in their standard state, the value of ΔH_f° is _____.
2. The standard heat of decomposition of a substance is determined by the following thermodynamic equation:



- a. Describe how the value of $\Delta H_{\text{decomp}}^\circ$ is being interpreted in the above decomposition.
- b. What is the correct sign for the enthalpy change in this decomposition reaction?

Write the correct answer in the space to the left. Show all your work in the space provided.

3. _____ $\Delta H_{\text{decomp}}^\circ$ of a substance is 20.4° kJ/mol. If a sample of 4.00 g of substance is heated until it decomposes, how much heat is absorbed?

4. _____ A substance has a specific heat of 0.001 J/g $^\circ\text{C}$. If a sample weighing 10.0 g is heated from 20°C to 100°C , how much heat is absorbed?

5. The product is a molecular compound of 10.0 g/mol, and the standard heat of formation of this compound

- a. What is the value of ΔH_f° for this reaction?

Answers For Reaction Energy Section 1

Rachel Sandford



Answers For Reaction Energy Section 1:

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