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web a derivatization that achieves this goes as $\text{molarity} \times \text{volume of solution in liter} = \text{moles}$ $\text{molarity} \times \text{volume in liter} = \text{mass moles} \times \text{molar mass}$ numerical answers 1

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web since the mole fraction of the O_2 is 0.050 it contributes 0.050 or if you will 5.0% of the total pressure thus total pressure is 0.2105042 atm problem 9 a sample of 1.43 g of helium and an unweighed quantity of O_2 are mixed in a flask at room temperature