

CHAPTER 10: Solutions

Reaction Energy

SECTION 1

Write a balanced chemical equation for the following reactions in the space provided.

- For elements in their standard state, the value of ΔH_f° is _____.
- The formation and decomposition of a substance is represented by the following thermodynamic equations:



- Is the energy being added to or is it being subtracted from the decomposition?
- What is the correct sign for the enthalpy change in this decomposition reaction?

Write the correct answer in the space to the left. Show all your work in the space provided.

- _____ $\Delta H_{\text{f}}^\circ$ of a substance is 207 kJ/mol . Calculate the $\Delta H_{\text{d}}^\circ$ of the decomposition of 1 mole of this substance.

- _____ A substance has a specific heat of $0.001 \text{ J/g}^\circ\text{C}$. If the specific heat capacity of a material increases by a factor of 10, the $\Delta H_{\text{f}}^\circ$ of the substance will be _____.

- The product is a molecular compound and the reactant is a ionic compound, and the reaction is exothermic. Is this correct?

- What is the value of $\Delta H_{\text{f}}^\circ$ for this reaction?

Answers For Reaction Energy Section 1

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