

Lab 10: Calculations with a Chemical Structure

Objectives:

1. To derive the structural formula of a given chemical compound without performing stoichiometric calculations.
2. To calculate the number of moles of a solid from weighing information in the lab.
3. To determine the molar mass of a solid.
4. To determine the stoichiometric amount of a solid compound that would be produced.
5. To compare the theoretical amount to the actual amount of substance formed and calculate the percent yield.

Reaction Information:

Reaction of Sodium Chloride and Silver Nitrate: $NaCl$ and $AgNO_3$

Reaction of Silver Chloride and Sodium Nitrate: $AgCl$ and $NaNO_3$

Reaction of Silver Chloride and Sodium Nitrate: $AgCl$ and $NaNO_3$

Reaction of Silver Chloride and Sodium Nitrate: $AgCl$ and $NaNO_3$

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Reaction of Silver Chloride and Sodium Nitrate: $AgCl$ and $NaNO_3$

Procedure:

Reaction and Calculations:

1. Reaction:

$$\text{moles of sodium chloride} = \frac{\text{mass of NaCl}}{\text{molar mass of NaCl}} = \frac{0.5 \text{ g}}{58.44 \text{ g/mol}} = 0.00855 \text{ mol}$$

$$\text{moles of silver chloride} = \frac{\text{mass of AgCl}}{\text{molar mass of AgCl}} = \frac{0.5 \text{ g}}{143.32 \text{ g/mol}} = 0.00349 \text{ mol}$$

$$\text{moles of sodium nitrate produced} = \text{moles of AgCl} = 0.00349 \text{ mol}$$

$$\text{mass of sodium nitrate produced} = \text{moles of NaNO}_3 \times \text{molar mass of NaNO}_3 = 0.00349 \text{ mol} \times 85.00 \text{ g/mol} = 0.297 \text{ g}$$

2. Reaction: $NaCl + AgNO_3 \rightarrow AgCl + NaNO_3$

3. Reaction:

$$\text{mass of NaCl} = \text{mass of AgCl} \times \frac{\text{molar mass of NaCl}}{\text{molar mass of AgCl}} = 0.5 \text{ g} \times \frac{58.44 \text{ g/mol}}{143.32 \text{ g/mol}} = 0.204 \text{ g}$$

$$\text{mass of NaCl} = \text{mass of AgCl} \times \frac{\text{molar mass of NaCl}}{\text{molar mass of AgCl}} = 0.5 \text{ g} \times \frac{58.44 \text{ g/mol}}{143.32 \text{ g/mol}} = 0.204 \text{ g}$$

$$\text{mass of NaCl} = \text{mass of AgCl} \times \frac{\text{molar mass of NaCl}}{\text{molar mass of AgCl}} = 0.5 \text{ g} \times \frac{58.44 \text{ g/mol}}{143.32 \text{ g/mol}} = 0.204 \text{ g}$$

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