

QUESTION 2 (contd)

⑤. (ii) The system is an exothermic one, when it is cooled, the equilibrium will adjust to annul the cooling effect by releasing more heat. And this favours the forward reaction.

$$\text{⑥. (i)} \quad \text{PbCl}_2 \longrightarrow \text{Pb}^{2+} + 2\text{Cl}^-$$

$$\text{No. of moles of PbCl}_2 = \frac{\text{mass}}{\text{molar mass}} = \frac{92.7}{207 + (35.5 \times 2)}$$

$$= \frac{92.7}{278}$$

$$= 0.33345 \text{ moles}$$

$$\text{No. of moles of Pb}^{2+} \text{ ionized} = 0.33345 \text{ moles}$$

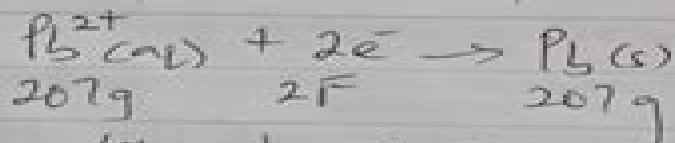
$$\text{mass of lead present} = \text{no. of moles} \times \text{molar mass}$$

$$= 0.33345 \times 207$$

$$= 69.0 \text{ grams}$$

$$\text{⑥. (ii)} \quad \text{Quantity of electricity used} = It = I \times 4.8 \times 3600$$

$$= (17280 \text{ I}) \text{ C}$$



2x 96500 liberate 207 grams of Pb

$$17280 \text{ I will liberate } \frac{17280 \text{ I} \times 207}{2 \times 96500} = 207$$

$$I = \frac{2 \times 96500 \times 207}{17280 \times 207} = 11.17 \text{ Amperes}$$

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 - The Rise of Digital Reading Answers Of Chemistry 2 Essay And 1 Obj
 - Advantages of eBooks Over Traditional Books
2. Identifying Answers Of Chemistry 2 Essay And 1 Obj
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Answers Of Chemistry 2 Essay And 1 Obj
 - User-Friendly Interface
4. Exploring eBook Recommendations from Answers Of Chemistry 2 Essay And 1 Obj
 - Personalized Recommendations
 - Answers Of Chemistry 2 Essay And 1 Obj User Reviews and Ratings
 - Answers Of Chemistry 2 Essay And 1 Obj and Bestseller Lists
5. Accessing Answers Of Chemistry 2 Essay And 1 Obj Free and Paid eBooks
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 - Answers Of Chemistry 2 Essay And 1 Obj eBook Subscription Services
 - Answers Of Chemistry 2 Essay And 1 Obj Budget-Friendly Options
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 - Answers Of Chemistry 2 Essay And 1 Obj Enhanced eBook Features
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 - Highlighting and Note-Taking Answers Of Chemistry 2 Essay And 1 Obj
 - Interactive Elements Answers Of Chemistry 2 Essay And 1 Obj
 8. Staying Engaged with Answers Of Chemistry 2 Essay And 1 Obj
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Answers Of Chemistry 2 Essay And 1 Obj
 9. Balancing eBooks and Physical Books Answers Of Chemistry 2 Essay And 1 Obj
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Answers Of Chemistry 2 Essay And 1 Obj
 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
 11. Cultivating a Reading Routine Answers Of Chemistry 2 Essay And 1 Obj
 - Setting Reading Goals Answers Of Chemistry 2 Essay And 1 Obj
 - Carving Out Dedicated Reading Time
 12. Sourcing Reliable Information of Answers Of Chemistry 2 Essay And 1 Obj
 - Fact-Checking eBook Content of Answers Of Chemistry 2 Essay And 1 Obj
 - Distinguishing Credible Sources
 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
 14. Embracing eBook Trends
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