# ACS ANALYTICAL CHEMISTRY STUDY GUIDE

- kf: kf=kf\*a
- EDTA titations: 1. before eq EDTA binds to metal 1:1

2. At eq we assume that all metal ions can bind to EDTA and then use kf

 Potentiometry: an analytical method in which an electric potential difference (avoltage) of a cell is measured

- Reducing agent: Donates electrons and becomes oxidized.
- oxidizing agent: Accepts electrons and becomes reduced.
- coulametry: science of measuring change and relating it to analyte
- cathode: the electrode at which reduction occurs
- Anode: the electrode at which oxidation occurs
- Ecell: Ecell=Ered-Eox
- Faraday's Equation: E=E-(.o592/n)log(a/b)
- ||: anode always written on left of ||

 ampeometry: elec techniques where the cell current is measured where pot diffbetween the indicator and referance electrode is controlled

- Clark electrode: Electrode for oxygen
- Biosensors: Detecting devices that can detect specific substances in cells
- amp chromo det: point, size, or charge
- limiting current: the diff between current and steady state max
- · Anodic vol: Very sensative for analyzing metals into ppb
- · fluroescence: used to measure singlets, atoms like rigid aromatic rings, highsensativity
- · phospherecence: involves emission of light trip- singlet, much lower than fluro
- Transmittance: A=-log(T) or 2-log(T%)
- atomic absorbtion: lamp-flame-mono-detector
- atomic emission: flame-mono-detector
- AA hollow cathode lamp: used to form very narrow band lights
- AA atomizer: Heat sample to very high temps to break all chemical bonds, hotenough to

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**KJ Lindholm-Leary** 

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