

Limiting Reagent & Percent Yield Practice Worksheet

1. When copper (II) chloride reacts with sodium nitrate, copper (II) nitrate and sodium chloride are formed.

a. Write the balanced equation for the reaction given above:



- b. If 15 grams of copper (II) chloride react with 20 grams of sodium nitrate, how much sodium chloride can be formed?

$$15 \text{ g CuCl}_2 \times \frac{1 \text{ mol CuCl}_2}{134.362 \text{ g CuCl}_2} \times \frac{2 \text{ mol NaCl}}{1 \text{ mol CuCl}_2} \times \frac{58.443 \text{ g NaCl}}{1 \text{ mol NaCl}} = 13 \text{ g NaCl}$$

$$20 \text{ g NaNO}_3 \times \frac{1 \text{ mol NaNO}_3}{84.997 \text{ g NaNO}_3} \times \frac{2 \text{ mol NaCl}}{2 \text{ mol NaNO}_3} \times \frac{58.443 \text{ g NaCl}}{1 \text{ mol NaCl}} = 13.8 \text{ g NaCl}$$

Only 13 g can be formed

- c. What is the limiting reagent for the reaction in #2? copper (II) chloride

- d. How many grams of copper (II) nitrate is formed?

$$15 \text{ g CuCl}_2 \times \frac{1 \text{ mol CuCl}_2}{134.452 \text{ g CuCl}_2} \times \frac{1 \text{ mol Cu(NO}_3)_2}{1 \text{ mol CuCl}_2} \times \frac{187.467 \text{ g Cu(NO}_3)_2}{1 \text{ mol Cu(NO}_3)_2} = 20.9 \text{ g Cu(NO}_3)_2$$

- e. How much of the excess reagent is left over in this reaction?

$$15 \text{ g CuCl}_2 \times \frac{1 \text{ mol CuCl}_2}{134.452 \text{ g CuCl}_2} \times \frac{2 \text{ mol NaNO}_3}{1 \text{ mol CuCl}_2} \times \frac{84.997 \text{ g NaNO}_3}{1 \text{ mol NaNO}_3} = 19 \text{ g NaNO}_3$$

$$20 \text{ g NaNO}_3 - 19 \text{ g NaNO}_3 = 1 \text{ g NaNO}_3$$

- f. If 11.3 grams of sodium chloride are formed in the reaction described in problem #2, what is the percent yield of this reaction?

$$\text{Percent yield} = \frac{\text{actual yield}}{\text{theoretical yield}} \times 100 \text{ percent}$$

$$11.3/13.0 \times 100\% = 86.9\%$$

2. When lead (II) nitrate reacts with sodium iodide, sodium nitrate and lead (II) iodide are formed.

a. Balance the following equation:



- b. If I start with 25.0 grams of lead (II) nitrate and 15.0 grams of sodium iodide, how many grams of sodium nitrate can be formed?

$$25.0 \text{ g Pb(NO}_3)_2 \times \frac{1 \text{ mol Pb(NO}_3)_2}{331.214 \text{ g Pb(NO}_3)_2} \times \frac{2 \text{ mol NaNO}_3}{1 \text{ mol Pb(NO}_3)_2} \times \frac{84.997 \text{ g NaNO}_3}{1 \text{ mol NaNO}_3} = 12.8 \text{ g NaNO}_3$$

$$15.0 \text{ g NaI} \times \frac{1 \text{ mol NaI}}{149.89 \text{ g NaI}} \times \frac{2 \text{ mol NaNO}_3}{2 \text{ mol NaI}} \times \frac{84.997 \text{ g NaNO}_3}{1 \text{ mol NaNO}_3} = 8.51 \text{ g NaNO}_3$$

8.51 g NaNO₃ can be made

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J Elliott



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1. Understanding the eBook 20cavalcade Publishing Limiting Reagent Worksheet
 - The Rise of Digital Reading 20cavalcade Publishing Limiting Reagent Worksheet
 - Advantages of eBooks Over Traditional Books
2. Identifying 20cavalcade Publishing Limiting Reagent Worksheet
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an 20cavalcade Publishing Limiting Reagent Worksheet
 - User-Friendly Interface
4. Exploring eBook Recommendations from 20cavalcade Publishing Limiting Reagent Worksheet
 - Personalized Recommendations
 - 20cavalcade Publishing Limiting Reagent Worksheet User Reviews and Ratings
 - 20cavalcade Publishing Limiting Reagent Worksheet and Bestseller Lists
5. Accessing 20cavalcade Publishing Limiting Reagent Worksheet Free and Paid eBooks
 - 20cavalcade Publishing Limiting Reagent Worksheet Public Domain eBooks
 - 20cavalcade Publishing Limiting Reagent Worksheet eBook Subscription Services
 - 20cavalcade Publishing Limiting Reagent Worksheet Budget-Friendly Options
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- ePub, PDF, MOBI, and More
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 - 20cavalcade Publishing Limiting Reagent Worksheet Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of 20cavalcade Publishing Limiting Reagent Worksheet
 - Highlighting and Note-Taking 20cavalcade Publishing Limiting Reagent Worksheet
 - Interactive Elements 20cavalcade Publishing Limiting Reagent Worksheet
 8. Staying Engaged with 20cavalcade Publishing Limiting Reagent Worksheet
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers 20cavalcade Publishing Limiting Reagent Worksheet
 9. Balancing eBooks and Physical Books 20cavalcade Publishing Limiting Reagent Worksheet
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection 20cavalcade Publishing Limiting Reagent Worksheet
 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
 11. Cultivating a Reading Routine 20cavalcade Publishing Limiting Reagent Worksheet
 - Setting Reading Goals 20cavalcade Publishing Limiting Reagent Worksheet
 - Carving Out Dedicated Reading Time
 12. Sourcing Reliable Information of 20cavalcade Publishing Limiting Reagent Worksheet
 - Fact-Checking eBook Content of 20cavalcade Publishing Limiting Reagent Worksheet
 - Distinguishing Credible Sources
 13. Promoting Lifelong Learning
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