

**Review for the Full Year General Chemistry
American Chemical Society Exam**

(a) Thermodynamics, (b) Transition Metals, (c) Redox and Electrochemistry, (d) Nuclear and (e) Organic

THERMODYNAMICS

84. A 10.0 g sample of silver is heated to 100.0 °C and then added to 20.0 g of water at 23.0 °C in an insulated calorimeter. At thermal equilibrium the temperature of the system was measured as 25.0 °C. What is the specific heat of silver? [specific heat = $4.2 \text{ J g}^{-1} \text{ }^{\circ}\text{C}^{-1}$]

- (A) 0.11 J/g K (C) 17 J/g K
(B) 0.22 J/g K (D) 34 J/g K

85. What is ΔH in kJ for the reaction:



- (A) -1830 kJ (B) -2284 kJ
(C) -2510 kJ (D) -1605 kJ

86. Given these equations:



calculate the heat of formation of $\text{SO}_3(\text{g})$.

- (A) -500 kJ mol⁻¹ (C) +100 kJ mol⁻¹
(B) -400 kJ mol⁻¹ (D) +200 kJ mol⁻¹

87. Given:



How much energy is produced in the combustion of 28 g of CO? [Atomic Masses: C 12.0 g mol⁻¹; O 16.0 g mol⁻¹]

- (A) 88 kJ (B) 109 kJ
(C) 175 kJ (D) 284 kJ

88. Calculate the value of ΔH (in kJ mol⁻¹) for the reaction.



Bond Energies (kJ mol⁻¹)

H-H	435
N=N	945 (in N ₂)
N-H	389

- (A) 2340 kJ of heat absorbed
(B) 213 kJ of heat absorbed
(C) 2340 kJ of heat evolved
(D) 83 kJ of heat evolved

89. Which change is likely to be accompanied by the greatest increase in entropy?



90. For which process is the entropy change per mole the largest at constant temperature?



91. In which process is entropy decreased?

- (A) dissolving sugar in water
(B) expanding a gas
(C) evaporating a liquid
(D) freezing water

92. When $\text{Al}_2\text{O}_3(\text{s})$ is formed from the elements at standard conditions, the values of ΔH° and ΔG° at 298 K are -1676 kJ mol⁻¹ and -1577 kJ mol⁻¹, respectively. The standard entropy of formation per mole, in joules per degree, will be

- (A) -332 (B) -157 (C) -93.3
(D) -0.0933 (E) +15.7

93. Vaporization of a liquid is an example of a process for which

- (A) ΔH , ΔS , and ΔG are positive at all temperatures.
(B) ΔH and ΔS are positive.
(C) ΔG is negative at low temperatures, positive at high temperatures.
(D) $\Delta H = \Delta S$

94. A particular chemical reaction has a negative ΔH and negative ΔS . Which statement is correct?

- (A) The reaction is spontaneous at all temperatures.
(B) The reaction is nonspontaneous at all temperatures.
(C) The reaction becomes spontaneous as temperature increases.
(D) The reaction becomes spontaneous as temperature decreases.

American Chemical Society General Chemistry 1 Exam

Sheila Tobias



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thinking The instructor's role is to facilitate the development of student concepts and process skills not to simply deliver content to the students The first part of this book introduces the theoretical and philosophical foundations of POGIL pedagogy and summarizes the literature demonstrating its efficacy The second part of the book focusses on implementing POGIL covering the formation and effective management of student teams offering guidance on the selection and writing of POGIL activities as well as on facilitation teaching large classes and assessment The book concludes with examples of implementation in STEM and non STEM disciplines as well as guidance on how to get started Appendices provide additional resources and information about The POGIL Project

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