#### Unit 08 - Stoichiometry - Worksheet 1

Answer the following questions in the space provided at the left.

18 mol O<sub>2</sub> 1. How many moles of oxygen are produced from twelve moles of Potassium chlorate?

2KCLO<sub>3</sub> → 2KCl + 3O<sub>2</sub>

12 mol KClO<sub>3</sub> 3 mol O<sub>2</sub> 2 mol KClO<sub>3</sub>

21 mol O<sub>2</sub> 2. How many moles of oxygen are produced from 14 moles of Potassium chlorate?

2KCLO<sub>3</sub> → 2KCl + 3O<sub>2</sub>

14 mol KetO<sub>3</sub> 3 mol O<sub>2</sub> 2 mol KClO<sub>3</sub>

\_\_45 mol H<sub>2</sub>O\_\_\_ 3. How many moles of water are produced from 22.5 moles of oxygen?
2H<sub>2</sub> + O<sub>1</sub>Z → 2H<sub>2</sub>O

22.5 mol-O<sub>2</sub> 2 mol H<sub>2</sub>O 1 mol-O<sub>2</sub>

\_20 mol HCI\_\_\_\_ 4. Calculate the number of moles of Hydrogen chloride produced from 10 moles of hydrogen.

H<sub>2</sub> + Cl<sub>2</sub> → 2HCl

10 mol H<sub>2</sub> 2 mol HCl 1 mol H<sub>2</sub>

21 mol Cl<sub>2</sub>
5. Calculate the number of moles of chlorine needed to form 14 moles of Iron (III) chloride.

2Fe + 3Cl<sub>2</sub> → 2FeCl<sub>3</sub>

14 mol EeCl<sub>3</sub> 3 mol Cl<sub>2</sub> 2 mol FeCl<sub>3</sub>

# **Answer Sheet To Chemistry Quiz Review Stoichiometry**

Marcel A. Müller

#### **Answer Sheet To Chemistry Quiz Review Stoichiometry:**

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questions on Atoms atomic spectrum atomic absorption spectrum atomic emission spectrum molecules azimuthal quantum number Bohr s model Bohr s atomic model defects charge to mass ratio of electron discovery of electron discovery of neutron discovery of proton dual nature of matter electron charge electron distribution electron radius and energy derivation electron velocity electronic configuration of elements energy of revolving electron fundamental particles Heisenberg's uncertainty principle hydrogen spectrum magnetic quantum number mass of electron metallic crystals properties Moseley law neutron properties orbital concept photons wave number Planck's quantum theory properties of cathode rays properties of positive rays quantum numbers quantum theory Rutherford model of atom shapes of orbitals spin quantum number what is spectrum x rays and atomic number The Basic Chemistry MCQ PDF e Book Chapter 2 practice test to solve MCQ questions on Basic chemistry atomic mass atoms molecules Avogadro s law combustion analysis empirical formula isotopes mass spectrometer molar volume molecular ions moles positive and negative ions relative abundance spectrometer and stoichiometry The Chemical Bonding MCQ PDF e Book Chapter 3 practice test to solve MCQ questions on Chemical bonding chemical combinations atomic radii atomic radius periodic table atomic ionic and covalent radii atoms and molecules bond formation covalent radius electron affinity electronegativity electronegativity periodic table higher ionization energies ionic radius ionization energies ionization energy periodic table Lewis concept and modern periodic table The Experimental Techniques MCQ PDF e Book Chapter 4 practice test to solve MCQ questions on Experimental techniques chromatography crystallization filter paper filtration filtration crucibles solvent extraction and sublimation The Gases MCQ PDF e Book Chapter 5 practice test to solve MCQ questions on Gas laws gas properties kinetic molecular theory of gases ideal gas constant ideal gas density liquefaction of gases absolute zero derivation applications of Daltons law Avogadro s law Boyle s law Charles law Daltons law diffusion and effusion Graham s law of diffusion ideality deviations kinetic interpretation of temperature liquids properties non ideal behavior of gases partial pressure calculations plasma state pressure units solid s properties states of matter thermometry scales and van der Waals equation The Liquids and Solids MCQ PDF e Book Chapter 6 practice test to solve MCQ questions on Liquid crystals types of solids classification of solids comparison in solids covalent solids properties of crystalline solids Avogadro number determination boiling point external pressure boiling points crystal lattice crystals and classification cubic close packing diamond structure dipole dipole forces dipole induced dipole forces dynamic equilibrium energy changes intermolecular attractions hexagonal close packing hydrogen bonding intermolecular forces London dispersion forces metallic crystals properties metallic solids metal s structure molecular solids phase changes energies properties of covalent crystals solid iodine structure unit cell and vapor pressure Kaplan SAT Subject Test Chemistry 2015-2016 Kaplan Test Prep, 2015-03-03 Essential strategies practice and review to ace the SAT Subject Test Chemistry Getting into a top college has never been more difficult Students need to distinguish themselves from the crowd and scoring well on a SAT Subject Test gives students a competitive edge Kaplan's SAT Subject Test Chemistry is the most up to date

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