

CHAPTER 10: Reactions

Reaction Energy**SECTION 1**

Write a balanced chemical equation for the following reactions in the space provided.

- For elements in their standard state, the value of ΔH_f° is _____.
- The formation and decomposition of a substance is represented by the following thermodynamic equations:



- Is the energy being added to or is it being subtracted from the decomposition?
- What is the correct sign for the enthalpy change in this decomposition reaction?

Write the correct answer in the space to the left. Show all your work in the space provided.

- _____ $\Delta H_{\text{f}}^\circ$ for a reaction is 2074° kJ/mol of a single bond energy value would the bond dissociation be?
- _____ A compound has a specific heat of $0.001 \text{ J/g}^\circ\text{C}$. If the specific heat capacity is a constant amount for a compound and is $0.001 \text{ J/g}^\circ\text{C}$, calculate the bond dissociation energy for 100°C and 0.001°C .
- The product is a constant bond energy of $0.001 \text{ J/g}^\circ\text{C}$, and the constant bond energy is $0.001 \text{ J/g}^\circ\text{C}$.
 - What is the value of ΔH for this reaction?

Answers For Reaction Energy Section 1

Scott C. Dulebohn



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