## Stoichiometry: Mass - Mole Worksheet

Based on the following chemical equation, please answer the following questions!

Fe (SO<sub>4</sub>), + 6 NaCl --> 2 FeCl, + 3 Na<sub>2</sub>SO<sub>4</sub>

Calculate the amount of moles of NaCl needed to produce 75 g of FeCl,.

| 9 FeCl, | 1 mole FeCl, | mole NaCl |  |
|---------|--------------|-----------|--|
|         | 9 FeCl,      | mole FeCl |  |

Calculate the amount of Fe, (SO,), in grams produce 215 g Na, SO,.

| 9 Na <sub>3</sub> SO <sub>4</sub> | 1 mole Na SO                      | mole Fe <sub>2</sub> (SO <sub>4</sub> ), | 9 Fe <sub>2</sub> (SO <sub>4</sub> ),      |
|-----------------------------------|-----------------------------------|--|--|
|                                   | 9 Na <sub>2</sub> SO <sub>4</sub> | mole Na,SO,                              | 1 mole Fe <sub>2</sub> (SO <sub>4</sub> ), |

Calculate the amount in grams of FeCl, produced with 95g of Fe (SO<sub>4</sub>),

| 9 Fe <sub>2</sub> (SO <sub>4</sub> ), | mole Fe2(SO4) | mole FeCl,                 | 9 FeCl,      |
|---------------------------------------|---------------|----------------------------|--------------|
|                                       | 9 Fe,(SO,),   | mole Fe(SO <sub>4</sub> ), | 1 mole FeCl, |

Calculate the amount of moles of Fe<sub>2</sub>(SO<sub>4</sub>) needed to react with 123.5 g of NaCl

| 9 NaCl | 1 mole NaCl | mole Fe(SO <sub>e</sub> ), |  |
|--------|-------------|----------------------------|--|
|        | 9 NaCl      | mole NaCl                  |  |

Based on the following chemical equation, please answer the following questions!

| 9 Ca(HCO <sub>y</sub> ) <sub>2</sub> | 1 mole Ca(HCO <sub>5</sub> ), | mole CaCO,     |
|--------------------------------------|-------------------------------|----------------|
|                                      | 9 Ca(HCO,),                   | mole Ca(HCO,), |

# **Answers For Study Guide Mass Moles**

**Susan McMurry** 

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