

## BASIC CHEMISTRY – Part 2

Name: Answer Key

Date: \_\_\_\_\_

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### PART 1: MIXTURES:

Fill in the blanks from the information provided on the presentation.

- A **mixture** is two or more substances **blended** together (**not bonded**).
- Each substance in a mixture keeps all of its **chemical** and **physical** properties.
- Mixtures, unlike compounds, can be **separated** into their individual substances by **physical** means.
- Example: **salt** and **water** can be separated by **evaporation**.
- A **solution** is a mixture of **two** or more substances, one **dissolved** in another.
- Two parts of a solution are the **solute** and the **solvent**.
- **Solute** – substance being **dissolved**  
Examples: **sugar, salt**
- **Solvent** – substance doing the **dissolving**  
Example: **water**
- A solution becomes **saturated** when the solvent **holds** all the solute that it can. It can **no longer dissolve** solute.
- **Suspensions** are a mixture of particles **scattered** throughout another. The particles **do not** dissolve.  
Examples: **clouds, dust, fog, smoke, mud, blood**

#### Reflect:

1. Give another example of solutions including the solute and solvent.  
**Vinegar and Calcium Carbonate, etc...**
2. Can solutions be separated by chemical means?  
**Perhaps, but will still involve physical separation as well**

### PART 2: COMPOUNDS

- **Compounds** are **two** or **more** elements **chemically** joined by **bonds**.
- They are represented by chemical **formulas**.
- **Properties** of compounds are usually **different** than the atoms.  
Example: Hydrogen(**gas**) + Oxygen(**gas**) = Water (**liquid**)  
Example: Iron + Oxygen = **Rust**
- Compounds can be **separated** into their elements only if the chemical **bonds** are **broken**. This can be done by **chemical** means (chemical **reaction**)
- A **molecule** is the **smallest** part of a compound—like water—that has **all** the **properties** of the compound.

# Answer To Living By Chemistry Lessons

**M Walker**



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### **Table of Contents Answer To Living By Chemistry Lessons**

1. Understanding the eBook Answer To Living By Chemistry Lessons
  - The Rise of Digital Reading Answer To Living By Chemistry Lessons
  - Advantages of eBooks Over Traditional Books
2. Identifying Answer To Living By Chemistry Lessons
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Answer To Living By Chemistry Lessons
  - User-Friendly Interface
4. Exploring eBook Recommendations from Answer To Living By Chemistry Lessons
  - Personalized Recommendations
  - Answer To Living By Chemistry Lessons User Reviews and Ratings
  - Answer To Living By Chemistry Lessons and Bestseller Lists

5. Accessing Answer To Living By Chemistry Lessons Free and Paid eBooks
  - Answer To Living By Chemistry Lessons Public Domain eBooks
  - Answer To Living By Chemistry Lessons eBook Subscription Services
  - Answer To Living By Chemistry Lessons Budget-Friendly Options
6. Navigating Answer To Living By Chemistry Lessons eBook Formats
  - ePub, PDF, MOBI, and More
  - Answer To Living By Chemistry Lessons Compatibility with Devices
  - Answer To Living By Chemistry Lessons Enhanced eBook Features
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  - Adjustable Fonts and Text Sizes of Answer To Living By Chemistry Lessons
  - Highlighting and Note-Taking Answer To Living By Chemistry Lessons
  - Interactive Elements Answer To Living By Chemistry Lessons
8. Staying Engaged with Answer To Living By Chemistry Lessons
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Answer To Living By Chemistry Lessons
9. Balancing eBooks and Physical Books Answer To Living By Chemistry Lessons
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Answer To Living By Chemistry Lessons
10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
11. Cultivating a Reading Routine Answer To Living By Chemistry Lessons
  - Setting Reading Goals Answer To Living By Chemistry Lessons
  - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Answer To Living By Chemistry Lessons
  - Fact-Checking eBook Content of Answer To Living By Chemistry Lessons
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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