

# Chemical equilibrium (reversible reactions)

Comprehension Passages with Activities & Diverse Assessments.

## Reading Comprehension - 9

### Industrial Applications

Chemical equilibrium governs industrial reversible reactions optimizing production. The Haber process ( $N_2 + 3H_2 \rightleftharpoons 2NH_3$ ) uses high pressure (200 atm) and moderate temperature (400°C) with iron catalysts to maximize ammonia fertilizers, balancing yield (favored by low temperature) - rate (needing heat) in sulfuric acid synthesis ( $2SO_2 + O_2 \rightleftharpoons 2SO_3$ ). High pressure and  $V_2O_5$  catalysts favor  $SO_3$ , 1 converted to  $H_2SO_4$ . Equilibrium constants guide conditions: large  $K$  reactions (e.g.,  $SO_3$  formation) need less or removing products like  $NaCl$  shifts equilibrium right, boost output. These principles, rooted since the early 20th century, rely on Le Chatelier to tweak variables, ensuring efficiency & cost-effectiveness. Equilibrium's predictability drives chem engineering from fuels to pharmaceuticals, linking their global industry.

## Comprehension Questions

1. What is chemical equilibrium in reversible reactions?

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2. Why is equilibrium considered dynamic?

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3. How is the equilibrium constant ( $K_c$ ) calculated?

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4. How does Le Chatelier's Principle predict shifts?

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5. How does temperature affect equilibrium?

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## True or False

Read each statement and circle if it is True or False.

|                                             |   |   |
|---------------------------------------------|---|---|
| Minimum pressure reactions stop.            | ✓ | ✗ |
| Minimum is dynamic.                         | ✓ | ✗ |
| Low concentration favors.                   | ✓ | ✗ |
| Catalysts shift equilibrium with catalysts. | ✓ | ✗ |
| Temperature favors exothermic reactions.    | ✓ | ✗ |
| High pressure favors more gas moles.        | ✓ | ✗ |
| High reactions shift to products.           | ✓ | ✗ |
| Adding change equilibrium position.         | ✓ | ✗ |
| Initial pressure makes $K_c$ .              | ✓ | ✗ |
| High $K_c$ means $SO_3$ is favored.         | ✓ | ✗ |

Print & Digital Ready

# Answer Key 18 3 Reversible Reactions And Equilibrium

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### **Answer Key 18 3 Reversible Reactions And Equilibrium:**

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